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Effective Conversion of Metmyoglobin to Oxymyoglobin by Cysteine-Substituted Polyphenols

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Running Head: Reducing-effects of cysteinyl polyphenols on metmyoglobin

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ABSTRACT: Reaction products from the peroxidase-catalyzed oxidation of polyphenols in the presence of cysteine showed a potent activity for reducing metmyoglobin (MetMb) to bright-colored oxymyoglobin (MbO₂). High-performance liquid chromatography (HPLC) purification of the reaction products from catechin, chlorogenic acid, dihydrocaffeic acid, hydroxytyrosol, nordihydroguaiaretic acid, and rosmarinic acid afforded corresponding *S*-cysteinyl compounds, whose structures were determined by nuclear magnetic resonance (NMR) spectroscopy and mass spectrometry (MS). The isolated cysteinyl polyphenols showed a concentration-dependent reducing activity for MetMb to MbO₂ for initial one hour. However, after one hour, some of them decreased the amount of MbO₂ produced. The effect of the number of cysteinyl sulfur substitutions in polyphenols on both MetMb reduction and MbO₂ maintenance was examined using hydroxytyrosols with different numbers of cysteine substitutions; these hydroxytyrosols were synthesized from hydroxytyrosol and an *N*-acetylcysteine methyl ester. The hydroxytyrosol derivative that substituted with two *N*-acetylcysteine esters exhibited the most effective reducing activity without any effect on MbO₂.

KEYWORDS: *metmyoglobin reducing activity, cysteinyl polyphenols, peroxidase-catalyzed oxidation, bright meat color preservation*

INTRODUCTION

The bright red color of meat is typically an indication of its freshness. The red color is the result of oxymyoglobin (MbO_2), which is formed during the blooming process of meat when molecular oxygen reacts with the heme protein, myoglobin. It is well known that myoglobin exists in various redox states, depending on the state of the chelated Fe ion. Myoglobin is rapidly converted to MbO_2 in an oxygen atmosphere; however, the produced MbO_2 is not very stable and is readily oxidized to a brown-colored metmyoglobin (MetMb). This browning reduces the market value of fresh meat.¹ Therefore, the preservation of the bright red color of meat and meat products is very important for the food industry.²

Polyphenols are well-known bioactive constituents of plants and plant-derived foods. They have potent reducing properties, resulting in the inhibition of oxidative deterioration of food. We have previously attempted to utilize such polyphenols to prevent the oxidative browning of MbO_2 . However, the potent antioxidant polyphenols actually accelerated the oxidation of MbO_2 to MetMb.³ This effect could be attributed the pro-oxidant properties of these highly active polyphenols. As an alternative approach to obtain fresh red-colored meat, we attempted to reduce MetMb back to MbO_2 .⁴ The results showed that flavonol polyphenols such as quercetin and kaempferol

could partially convert MetMb back to MbO₂, and reduced production of the pro-oxidant quinone derivatives was identified as a possible mechanism. However, other polyphenols previously reported with a strong reducing potency could not produce sufficient amounts of MbO₂. We found that the addition of equimolar amounts of cysteine to the polyphenol solution was very effective.⁵ A main reason for this was that the pro-oxidant quinones, which were formed in the assay system, were quenched by the nucleophilic addition of the cysteinyl thiol. In another investigation, we also found that several cysteine-substituted polyphenols could enhance the antioxidant capacity against lipid oxidation.⁶ These results indicated that such cysteine-substituted polyphenols exhibited a useful enhancing property for the MetMb reducing and the maintenance of MbO₂, which is very important for preserving the color of fresh meat. In this investigation, we prepared various cysteine-substituted polyphenols and examined their MetMb reducing activity. We also investigated the effect of the number of *S*-cysteinyl-substitutions on the ability of the polyphenols to reduce MetMb using hydroxytyrosol.

MATERIALS AND METHODS

Chemicals and instruments. Myoglobin, gentisic acid (purity >98%),

protocatechuic acid (purity >97%), quercetin dihydrate (purity >95%), potassium hexacyanoferrate (III), sodium hydrosulfite, cysteine, and *N*-butoxycarbonylcysteine were obtained from Nacalai Tesque (Kyoto, Japan). *N*-Acetyl cysteine (purity >98%), caffeic acid (purity >98%), hydroxytyrosol (purity >98%), myricetin (purity >97%), morin (purity >90% as hydrate), and kaempferol (purity >97% as hydrate) were purchased from Tokyo Kasei (Tokyo, Japan). Catechin (purity >98% as hydrate), rosmarinic acid (purity >97%), chlorogenic acid (purity >95%), nordihydroguaiaretic acid (purity >97%), and dihydrocaffeic acid (purity >98%) were obtained from Sigma-Aldrich (St. Louis, MO). Taxifolin (purity >96%) was obtained from Funakoshi (Tokyo, Japan). Peroxidase (from horseradish) was purchased from Toyobo (Osaka, Japan). All organic solvents (extra pure grade or HPLC grade) were obtained from Nacalai Tesque. Distilled water produced by a water-distilling instrument, SA-2100E (EYELA, Tokyo, Japan), was used for all experiments. NMR spectra were obtained from an ECS-400 spectrometer (JEOL, Tokyo, Japan) or an AVANCE 300N spectrometer (Bruker BioSpin, Yokohama, Japan) using the manufacturer's pulse sequences for ^1H , ^{13}C , HH-COSY, NOESY, HMQC, and HMBC measurements. High-resolution mass spectrometric data were obtained using a XEVO QtofMS spectrometer (Waters Japan, Tokyo, Japan) in ESI mode and determination of each

molecular formula from the MS data was carried out using MassLynx v.4.1 software (Waters Japan). An LC-20AT low-pressure gradient system (Shimadzu, Kyoto, Japan) equipped with an SPD-M20A photodiode array detector (PDA) was employed for analytical HPLC. PDA data were analyzed using LabSolutions ver. 5.51 (Shimadzu). An LC-6AD system (Shimadzu) equipped with a UV-8011 detector (Tosoh, Tokyo, Japan) was used for preparative HPLC.

Preparation of MetMb. The MetMb solution was prepared from a stock solution of MbO₂, which was prepared using a method reported previously.⁴ Potassium hexacyanoferrate (III) (2 mg) was added to a solution of MbO₂ (120 μmol/L) in a phosphate buffer (50 mmol/L, pH 7.4, 3 mL) at 23 °C. After gentle stirring for 0.5 min, the solution was desalted three times at 14,000 g for 6 min, at 4 °C, with an Amicon Ultra-0.5, Ultracel-10 ultrafiltration membrane (Merck Millipore, Cork, Ireland) using phosphate buffer. The final residual solution (≈ 0.5 mL) was diluted with buffer, yielding 120 μmol/L of the MetMb solution (3 mL). The purity and concentration of the MetMb solution were confirmed using a UVmini 1240 UV/VIS spectrometer (Shimadzu).^{7,8}

Coupling Reaction between Polyphenols and Cysteine. To a straight vial of the solution of each polyphenol in methanol (18 mmol/L, 1 mL), cysteine in water (18

mmol/L, 1 mL), peroxidase in water (60 $\mu\text{mol/L}$, 20 μL), and H_2O_2 (400 mmol/L, 30 μL) were subsequently added and then the solution was made up to 4 mL with water. The solution was incubated at 25 $^\circ\text{C}$ for 8 min to afford oxidized polyphenol solution (4.5 mmol/L based on starting polyphenol concentration). Control experiments were carried out by similar procedure without addition of cysteine to make reference samples.

Measurement of MetMb reduction to MbO₂. Briefly, 50 mmol/L of phosphate buffer (pH 7.4, 250 μL), 120 $\mu\text{mol/L}$ of MetMb (150 μL ; in the same buffer), DMSO 10 μL , and the sample solution mentioned above (4.5 mmol/L, 40 μL) were subsequently added to a 96-well microplate. The microplate was then incubated at 37 $^\circ\text{C}$ in a Multiskan GO microplate reader (Thermo Fisher Scientific, Yokohama, Japan) for 1 h, and absorbance was measured at 500 and 582 nm. The MbO₂ and MetMb concentrations in the wells were calculated using the following equation⁵: MbO₂ ($\mu\text{mol/L}$) = $(89.7 \times A_{582}) - (32.9 \times A_{500})$ where A_{500} and A_{582} refer to absorbance at 500 nm and 582 nm, respectively. The reductive conversion efficiency from MetMb to MbO₂ was expressed as a percentage by calculating the MbO₂ concentration relative to the initial concentration (60 $\mu\text{mol/L}$) of MetMb.

Purification and Structure Identification of Cysteinyl Polyphenols.

Preparative HPLC conditions are summarized in a table in supporting information

section. HR-MS and NMR data of the following synthesized compounds are also summarized in the section.

Cysteinyl catechins. To a catechin (50 mg) solution in methanol (4 mL) and water (12 mL) were added cysteine (42 mg), peroxidase (4.6 mg), and H₂O₂ (0.4 mol/L, 360 μL) with stirring at 25 °C. After several minutes, methanol in the solution was removed *in vacuo* and the resulting solution was purified by preparative HPLC to give 2'-*S*-cysteinylcatechin (**1**) and 5'-*S*-cysteinylcatechin (**2**) in 16 and 22% yields, respectively.

Cysteinyl chlorogenic acid. A chlorogenic acid (50 mg) solution in methanol (3 mL) and water (10 mL), cysteine (103 mg), peroxidase (3.7 mg), and H₂O₂ (0.4 mol/L, 350 μL) were treated under similar conditions to those of catechin and the obtained reaction mixture was purified by preparative HPLC to give 2''-*S*-cysteinylchlorogenic acid (**3**) in 64% yield.

Cysteinyl dihydrocaffeic acids: A dihydrocaffeic acid (50 mg) solution in methanol (7 mL) and water (18 mL), cysteine (131 mg), peroxidase (6.6 mg), and H₂O₂ (0.4 mol/L, 680 μL) were treated under similar conditions to those of catechin and the obtained reaction mixture was purified by preparative HPLC to give 2',5'-bis(*S*-cysteinyl)dihydrocaffeic acid (**4**) and 5'-*S*-cysteinyl dihydrocaffeic acid (**5**) in

8% and 65% yields, respectively.

Cysteinyl hydroxytyrosols: A hydroxytyrosol (50 mg) solution in methanol (5 mL) and water (10 mL), cysteine (78 mg), peroxidase (8.1 mg), and H₂O₂ (0.4 mol/L, 350 μL) were treated under similar conditions to those of catechin and the obtained reaction mixture was purified by preparative HPLC to give 2',5'-bis(*S*-cysteinyl)hydroxytyrosol (**6**) and 5'-*S*-cysteinylhydroxytyrosol (**7**) in 5% and 46% yields, respectively.

Cysteinyl nordihydroguaiaretic acid: A nordihydroguaiaretic acid (50 mg) solution in methanol (6 mL) and water (10 mL), cysteine (40 mg), peroxidase (4.4 mg), and H₂O₂ (0.4 mol/L, 450 μL) were treated under similar conditions to those of catechin and the obtained reaction mixture was purified by preparative HPLC to give 5'-*S*-cysteinylnordihydroguaiaretic acid (**8**) in 20% yield.

Cysteinyl rosmarinic acid: A rosmarinic acid (50 mg) solution in methanol (3 mL) and water (10 mL), cysteine (101 mg), peroxidase (3.5mg), and H₂O₂ (0.4 mol/L, 350 μL) were treated under similar conditions to those of catechin and the obtained reaction mixture was purified by preparative HPLC to give 2'''-*S*-cysteinylrosmarinic acid (**9**) in 48% yield.

Preparation of 5'-*S*-(*N*-acetyl-*O*-methylcysteinyl)hydroxytyrosol (10**) and 2',5'(or 5',6')-*S*-bis[(*N*-acetyl-*O*-methyl)cysteinyl]hydroxytyrosol (**11**).** To a

solution of hydroxytyrosol (400 mg) and *N*-acetylcysteine methyl ester (551 mg) in acetonitrile (250 mL) was added slowly a fine powder of 2,2-diphenyl-1-picrylhydrazyl (DPPH, 2.1 g) with stirring at 23 °C. The solution was then evaporated *in vacuo* and filtered after addition of a small amount of the mixed solvent from ethyl acetate and acetone [6:1(v/v)]. The filtrate was subjected to silica gel column chromatography eluted with ethyl acetate-acetone (6:1) and then ethyl acetate-acetone (4:1) to give **10** (176 mg) and **11** (188 mg) in 21% and 15% yields, respectively. The HR-MS and NMR data of **10** and **11** are summarized in supporting information section.

Preparation of 2',5',6'- tris[*S*-(*N*-acetyl-*O*-methylcysteinyl)]hydroxytyrosol

(12) To a solution of di-*N*-acetyl-*O*-methylcysteine substituted hydroxytyrosol (**11**, 90 mg) and *N*-acetylcysteine methyl ester (40 mg) in acetone (50 mL) was added a fine powder of DPPH (140 mg) with stirring at 23 °C. After stirring for several minutes, the solution was evaporated and subjected to column chromatography on silica gel (Cosmosil 75-SR-II-PREP, Nacalai) eluted with ethyl acetate-acetone (1:2) to give **12** (60 mg) in 50% yield. The HR-MS and NMR data of **12** are summarized in supporting information section.

RESULTS AND DISCUSSION

Metmyoglobin-reducing activity of peroxide-catalyzed reaction products from polyphenols and cysteine. For the coupling of polyphenols with cysteine, we found that peroxidase-catalyzed oxidation gave the best result.⁹ Other methods that employ chemical reagents for this purpose either oxidized the cysteine first or required time-consuming follow-up to remove the reacted reagents. The MetMb-reducing activity of 40 μL of the peroxidase reaction solutions, which were prepared from equimolar concentrations of polyphenol and cysteine, was measured using a freshly prepared MetMb solution (60 $\mu\text{mol/L}$). The data, which are summarized in Figure 1, show that most reaction products exhibited remarkable MetMb-reducing activity, whereas the peroxidase-oxidized polyphenols without cysteine did not show such high activity, except for the products from kaempferol and morin. Among the polyphenols examined, hydroxytyrosol, catechin, and caffeic acid showed the highest efficiency, followed by kaempferol, morin, taxifolin, rosmarinic acid, nordihydroguaiaretic acid. We had previously reported a high-MetMb reducing activity of polyphenols in the presence of cysteine⁵ suggesting that cysteinyl polyphenols produced in the solutions might play a role in MetMb-reducing activity. We therefore isolated cysteinyl polyphenols from the reaction mixtures and measured their ability to convert MetMb reductively to MbO₂.

Isolation and structure identification of cysteinyl polyphenols. From the catechin reaction, two cysteinyl catechins (**1** and **2**) were detected using HPLC analysis and these were isolated using a preparative HPLC technique as described in the experimental section. High-resolution MS (HR-MS) of **1** and **2** revealed that they had the same molecular formula ($C_{18}H_{19}NO_3S$), indicating that both were isomers of mono-cysteine-substituted catechins. The substituted position was deduced by NMR data. The 1H NMR data of **1** showed a coupling constant that signified *ortho*-related protons in the B-ring, indicating that the cysteine was substituted at the 2''-position. C-H long-range correlations in HMBC of **1** were observed between C2' and H2, C6' and H2, and C2' and H2'', indicating that the substitution at position 2' was through cysteinyl sulfur linkage. **1** was therefore determined to be 2'-*S*-cysteinylcatechin (Figure 2). For the structure of **2**, a *meta*-related coupling constant (1.8 Hz) was observed between H2' and H6' of the 1H NMR data of B-ring. **2** was therefore determined to be 5'-*S*-cysteinylcatechin (Figure 2).

The peroxidase-catalyzed reaction of chlorogenic acid yielded a major coupling product (**3**), which had the molecular formula: $C_{19}H_{23}NO_{11}S$ as derived from HR-MS data. The position of the substituted cysteine was determined from 1H NMR data, which showed two *ortho*-coupling protons in a caffeate moiety. The structure of **3** was

therefore determined to be 2''-*S*-cysteinylchlorogenic acid (Figure 2).

Dihydrocaffeic acid gave major two coupling products (**4** and **5**) by the peroxidase-catalyzed reaction. These products were estimated to be di- and mono-cysteinyl dihydrocaffeic acid from the HR-MS data (m/z 419.0572 [M-H]⁻ and 300.0527 [M-H]⁻, respectively). The substituted cysteine in mono-cysteinyl dihydrocaffeic acid (**5**) was estimated to occupy the 5'-position, because the *meta*-related coupling constant (2.4 Hz) was clearly observed in doublet signals of H2' and H6'. **5** was therefore determined to be 5'-*S*-cysteinyl dihydrocaffeic acid. For the di-cysteine-substituted dihydrocaffeic acid (**4**), the ¹H NMR data were too complex to allow complete elucidation of the structure. An aromatic proton signal was observed at 7.01 ppm as a singlet and NOE was observed from the aromatic proton to H2 and H3 of the cysteine part. These data indicated that the positions of the two substituted cysteines were either 2' and 5' or 5' and 6'. Therefore, **4** was a di-cysteine-substituted dihydrocaffeic acid such as 2',5'-bis(*S*-cysteinyl) dihydrocaffeic acid or 5',6'-bis(*S*-cysteinyl) dihydrocaffeic acid (Figure 2).

A peroxidase-catalyzed reaction of hydroxytyrosol in the presence of cysteine yielded two products (**6** and **7**). HR-ESIMS analysis of both products revealed that **6** and **7** were di-cysteine-substituted hydroxytyrosol and mono-cysteine-substituted

hydroxytyrosol, respectively. The $^1\text{H-NMR}$ spectrum of **6** was very complex except for the observation of a singlet aromatic proton at 7.03 ppm. We synthesized bis[*S*-(*N*-butoxycarbonylcysteinyl)]hydroxytyrosol from hydroxytyrosol and *N*-butoxycarbonylcysteine. In the compound, clear NOE between the aromatic proton and H3 of the cysteine part was observed, indicating that the substituted positions of the cysteines were either 2' and 5' or 5' and 6'. Acid treatment of the di-(*N*-butoxycarbonyl)cysteinyl hydroxytyrosol gave **6**. Therefore, **6** should be 2',5' (or 5',6')-bis(*S*-cysteinyl)hydroxytyrosol. Mono-cysteine substituted hydroxytyrosol (**7**) was also analyzed by NMR. Two *meta*-coupled aromatic protons were clearly observed in the $^1\text{H NMR}$ of **7**, revealing that **7** was 5'-*S*-cysteinylhydroxytyrosol (Figure 2).

Nordihydroguaiaretic acid, a potent antioxidant of creosote bush, also gave a coupling product **8** from the peroxidase reaction. The substituted position of the cysteine was determined to be at the 5'-position of one of the aromatic rings, which was revealed from the *meta*-related coupling constant between H2' and H6' in the $^1\text{H NMR}$ of **8**. Thus, **8** was determined to be 5'-*S*-cysteinylnordihydroguaiaretic acid (Figure 2).

Rosmarinic acid, a labietae polyphenol, also yielded mono-cysteinyl rosmarinic acid (**9**) as a peroxidase-catalyzed reaction product. Rosmarinic acid has two tri-substituted benzene rings, however, the $^1\text{H NMR}$ data of **9** showed the existence of a

four-substituted benzene ring possessing two protons with *ortho* coupling constant (8.4 Hz). One of the protons of the benzene ring was relatively low-field shifted. Therefore, the substituted cysteine was determined to be at the 2-position of the benzene ring of a caffeoyl moiety of rosmarinic acid. Thus, **9** was 2'''-*S*-cysteinylrosmarinic acid (Figure 2).

Metmyoglobin-reducing activity of cysteinyl polyphenols. Figure 3 shows the percent ratio of MbO₂ produced in the reducing reaction of 60 μmol of MetMb by 60 – 600 μmol of cysteinyl polyphenols. All cysteinyl polyphenols examined showed remarkable and concentration-dependent reducing activity to MetMb for the first hour of the reaction. A decrease of MbO₂ was observed in the reaction with cysteinyl polyphenols **5**, **7**, **8**, and **9** after one hour. It should be noted that cysteinyl nordihydroguaiaretic acid (**8**) and cysteinyl rosmarinic acid (**9**) have two catechol structures and one of them remains in the intact form. We have reported that highly antioxidative polyphenols, most of which have the same catechol structure, enhance the oxidation of MbO₂ in part owing to their pro-oxidant property. Therefore, the intact catechol structures of **8** and **9** should enhance the oxidative decrease of the initially produced MbO₂. Cysteinyl polyphenols **5** (mono-cysteinyl dihydrocaffeic acid) and **7** (mono-cysteinyl hydroxytyrosol) have very similar structures. Although our previous

results showed that a mono-sulfur substitution reduces the pro-oxidant properties of highly reactive polyphenols^{4,5} it is possible that a mono-cysteine-substituted dihydrocaffeic acid and hydroxytyrosol still have potent reactivity.¹⁰ These polyphenols have a catechol structure with no conjugated electron-withdrawal groups, which have high reactivity and enhance both antioxidant and pro-oxidant activities. Therefore, the remaining pro-oxidative properties might exhibit oxidation of MbO₂ at the end of the one-hour reaction. Noticeably, di-cysteine-substituted dihydrocaffeic acid (**4**) and hydroxytyrosol (**6**) did not show obvious pro-oxidant activity, which supports the idea that multi-sulfur substitution might reduce the pro-oxidant activity to MbO₂, even in the presence of highly reactive polyphenols.

The effect of the number of sulfur substitutions in polyphenols on metmyoglobin reduction and oxymyoglobin oxidation. Our previous results showed that in cysteinyl dihydrocaffeic acids and hydroxytyrosols, the number of sulfur substitutions have different effects and up to three substituents can be made to the benzene region. The question then arose as to how many cysteine substitutions would exhibit the best results in the terms of MetMb reduction and MbO₂ production and retention. We attempted to synthesize tri-cysteine-substituted hydroxytyrosol, but were unsuccessful using the peroxide-catalyzed method as well as other chemical oxidation

reagents. The reason for the failure seemed to be faster oxidation of the thiol part of the cysteine. We have reported that oxidation of the thiol part of the cysteine can be prevented by protection of both α -amino and carboxylic groups.¹¹ Using *N*-acetylcysteine methyl ester (NAC-OCH₃) as such a protectant, we prepared multi-sulfur-substituted hydroxytyrosols. The oxidative coupling of hydroxytyrosol with NAC-OCH₃ was performed in anhydrous acetonitrile using DPPH (2,2-diphenyl-1-picrylhydrazyl) as an oxidation reagent. The reaction of hydroxytyrosol with 4 equivalents of DPPH in the presence of equimolar NAC-OCH₃ yielded mono-substituted hydroxytyrosol (**10**) and di-substituted hydroxytyrosol (**11**) in 21% and 15% yields, respectively. Additional DPPH oxidation of **11** with NAC-OCH₃ gave a tri-substituted hydroxytyrosol (**12**) in 50% yield (Figure 4). The MetMb-reducing activities of **10**, **11**, and **12** were measured and the data are illustrated in Figure 5. The effects of NAC-substituted hydroxytyrosols (**10**–**12**) and intact hydroxytyrosol on MbO₂ are presented in Figure 6. Hydroxytyrosol and mono-cysteinyl derivative **10** showed no MetMb-reducing activity, whereas the di-cysteinyl derivative **11** and the tri-cysteinyl derivative **12** showed strong and concentration-dependent reducing activity. In particular, a higher reducing efficiency was observed in **11** (Figure 5), revealing that the di-sulfur-substituted compounds were the most effective at the reducing MetMb. In

contrast, an additional substitution weakened the reducing activity of the di-substituted hydroxytyrosol. It should be noted that the reducing efficiency of the di-substituted hydroxytyrosol (**11**) reached almost a 100% for 3 h at 37 °C at a concentration of 600 $\mu\text{mol/L}$ versus 60 $\mu\text{mol/L}$ of myoglobin. However, mono-substituted hydroxytyrosol (**10**)-induced enhancement of MbO₂ oxidation was similar to that of hydroxytyrosol, while di- and tri-substituted hydroxytyrosols showed negligible enhancement (Figure 6). We have previously reported that the antioxidant capacity of a dihydrocaffeic ester was enhanced in the presence of a cysteine derivative.⁶ The antioxidation of dihydrocaffeic ester in the presence of a cysteine derivative in the production of mono-, di-, and tri-cysteinyl dihydrocaffeates. Furthermore, these mono- and di-substituted dihydrocaffeates mainly caused an increase in the length of the antioxidation period of dihydrocaffeate. Different reactivity to a peroxy radical was also observed among these cysteinyl dihydrocaffeates. These results clarified that the substitution of cysteine, and specifically a sulfur substitution, modulated the redox abilities of the original polyphenols. In the cysteinyl hydroxytyrosols, di-cysteine-substituted hydroxytyrosol (**11**) showed the most efficient reducing activity for MetMb and no prooxidant effect on MbO₂. It is well known that hydroxytyrosol is an antioxidant present in olives and is also the antioxidant part of oleuropein¹² which is the main polyphenol not only in young

olive fruits¹³ but also in olive leaves.¹⁴ A series of investigations, including the current study,³⁻⁵ have demonstrated that potent active polyphenols could be used for preservation of fresh meat color in the presence of cysteine or related thiol compounds. We have also shown that a simple peroxidase-catalyzed reaction can be used for the modification of polyphenols to effectively reduce brown-colored MetMb. This simple method, which requires catalytic amount of enzyme and minimum amounts of cysteine and hydrogen peroxide, is applicable to advanced utilization of the raw extracts of polyphenol-rich plant resources such as olive leaves, because the reaction solution may be used directly for fresh meat color preservation.

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Figure Captions

Figure 1. Effects of the products (600 $\mu\text{mol/L}$ based on starting polyphenols) from polyphenols in peroxidase-catalyzed oxidations in the presence (red bar) or absence (blue bar) of cysteine on MetMb reduction. Results are expressed as mean \pm standard deviation (SD) ($n=3$). Compounds **1**, 2'-*S*-cysteinylcatechin; **2**, 5'-*S*-cysteinylcatechin; **3**, 2''-*S*-cysteinylchlorogenic acid; **4**, 2',5'-bis(*S*-cysteinyl)dihydrocaffeic acid; **5**, 5'-*S*-cysteinyl-dihydrocaffeic acid; **6**, 2',5' (or 5',6')-bis(*S*-cysteinyl)hydroxytyrosol; **7**, 5'-*S*-cysteinylhydroxytyrosol; **8**, 5'-*S*-cysteinylnordihydroguaiaretic acid; **9**, 2'''-*S*-cysteinylrosmarinic acid.

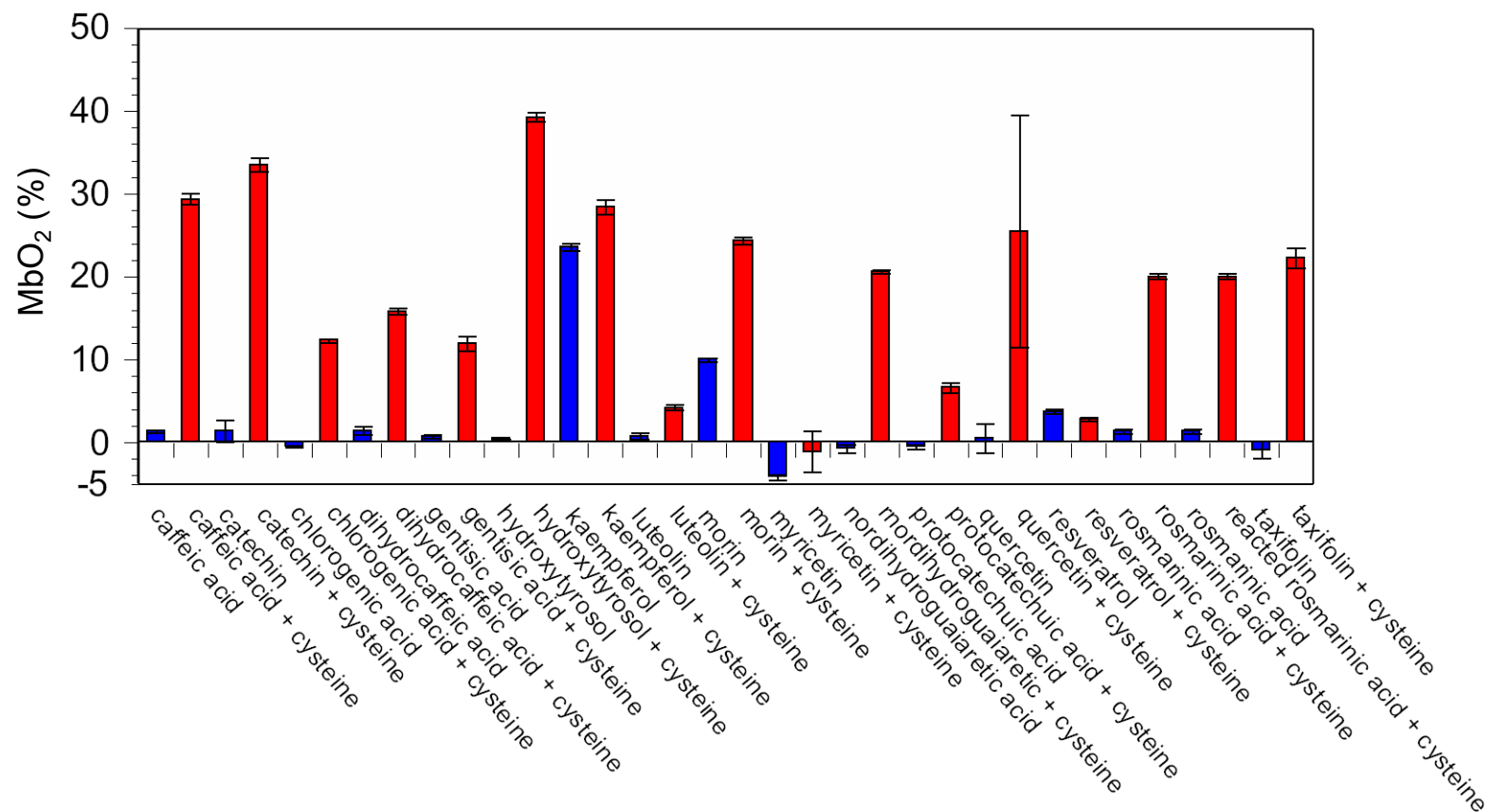
Figure 2. Chemical structures of the cysteinyl polyphenols identified in peroxidase reaction solutions.

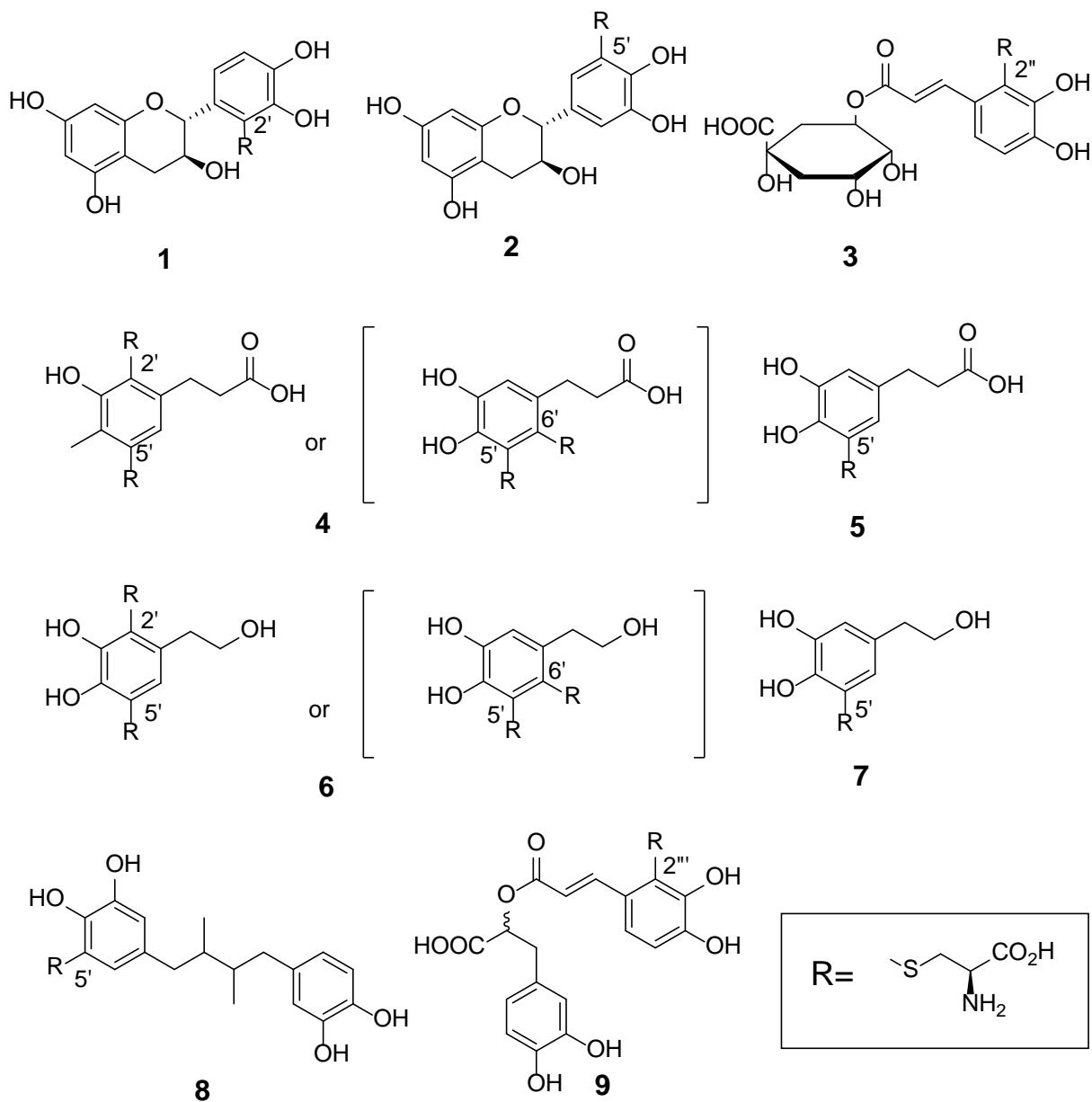
Figure 3. Effects of cysteinyl polyphenols (**1** – **9**) on MetMb reduction. All analytical experiments were performed in triplicate and the data were expressed as means (SDs were smaller than symbols).

Figure 4. Chemical structures of (*N*-acetyl-*O*-methyl)cysteinyl hydroxytyrosols.

Figure 5. Effects of (N-acetyl-O-methyl)cysteiny hydroxytyrosols (**10 – 12**) and hydroxytyrosol on MetMb reduction. All analytical experiments were performed in triplicate and the data were expressed as mean \pm standard deviation (most SDs were smaller than symbols).

Figure 6. Effects of (*N*-acetyl-*O*-methyl)cysteiny hydroxytyrosols (**10 – 12**) and intact hydroxytyrosol as a reference compound with no-cysteine substitution on MbO₂ maintenance. All analytical experiments were performed in triplicate and the data was expressed as mean \pm standard deviation (most SDs were smaller than symbols).





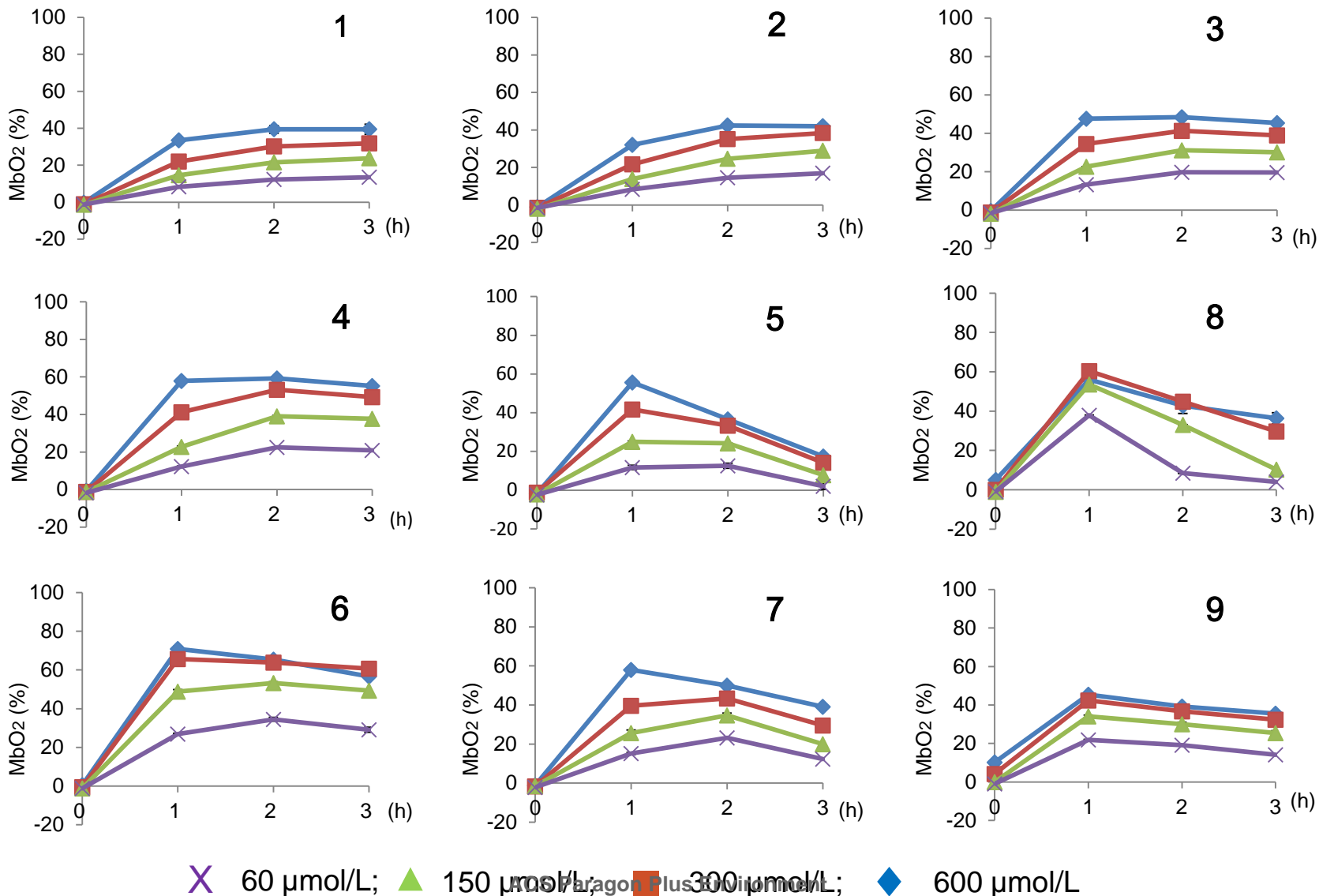


Fig.3

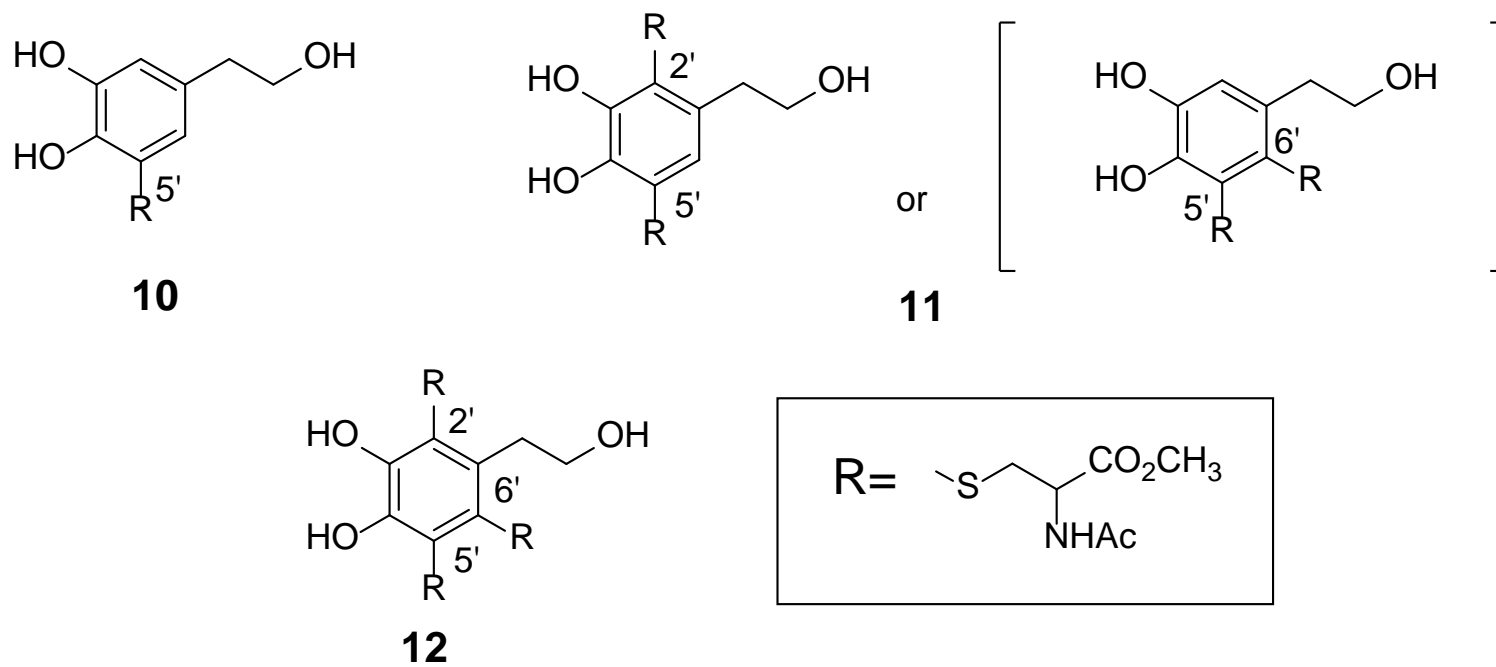


Fig. 4

