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Modulating the concentrations of reactive oxygen and nitrogen species and oxygen in water with helium and argon gas and plasma jets

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Herein, we employed UV-Vis spectroscopy to monitor real-time changes in the oxygen tension and concentration of reactive oxygen and nitrogen species (RONS) in deionized (DI) water during treatments with helium (He) and argon (Ar) gas and plasma jets. He and Ar gas jets are both shown to de-oxygenate DI water with He being more efficient than Ar, whilst the plasma jets deliver and regulate the concentrations of hydrogen peroxide (H₂O₂), nitrite (NO₂⁻) and nitrate (NO₃⁻) in deionized (DI) water. The H₂O₂ and NO₃⁻ production efficiency varied between He

and Ar plasma jets, but was similar for NO_2^- . Whilst DI water fully equilibrated with ambient air prior to treatment, was de-oxygenated by both plasma jets, when the DI water was first de-oxygenated by an inert gas jet treatment, both plasma jets were found to be capable of oxygenating the DI water. These insights were then used to show how different combinations of plasma jet and inert gas jet treatments can be used to modulate O_2 tension and RONS chemistry. Finally, potential further improvements to improve control in the use of plasma jets in regulating O_2 and RONS are discussed.

1. Introduction

Cold atmospheric helium (He) or argon (Ar) plasma jets (herein referred to as plasma jets) impinging on air, generate a cocktail of reactive oxygen species (ROS) and reactive nitrogen species (RNS), or collectively RONS.¹⁻⁵⁾ When these plasma jets are aimed at an aqueous liquid target, the RONS in the effluent of the flowing gas can dissolve into the liquid phase and from there participate in chemical reactions; e.g. protein oxidation⁶⁾ or redox reactions.⁷⁾ It was also shown that the inert gas of plasma jets can deoxygenate DI water.^{8,9)} In the case of applications to biological media, RONS species could potentially interact with biological molecules and/or components of the cells.⁷⁾ And cell fate is known to be sensitive to aqueous oxygen [$O_2(aq)$] tension.^{8,10)} We also note that there are many different designs of plasma sources,¹¹⁾ which will produce different RONS chemistry in aqueous liquids. In addition, the composition of the surrounding air such as oxygen and nitrogen content¹²⁾ and humidity,¹³⁾ as well as the initial liquid chemistry,¹⁴⁾ will significantly influence the resultant RONS chemistry in the plasma jet-treated liquid. Another important parameter is whether or not the plasma jet is in contact with the liquid surface, which can also significantly influence the resultant RONS chemistry in the liquid.¹⁵⁾

He and Ar are commonly chosen as feed gases because plasma jets formed from these gases can be operated with relatively low breakdown thresholds and sustainable voltages,¹⁶⁾ and because of the relative affordability of He and Ar gases. In order to achieve control, efficiency, and predictability over processing outcomes, it is important to understand the exact nature of He and Ar plasma jets/water interactions and what process variables most influence these. This is particularly important where He and Ar plasma jets exhibit different discharge characteristics¹⁷⁻²⁰⁾ which could potentially influence the delivery of RONS into the target liquid. In order to measure the real-time changes in the concentrations of $O_2(aq)$ and RONS species in DI water during the plasma jet treatment, we developed a UV-Vis spectroscopy (UV-Vis) method.^{9,21-25)} We have shown in a previous study, UV-Vis spectroscopy can be used to accurately determine the concentrations of H_2O_2 , NO_2^- and NO_3^- in reference solutions, with mixed concentrations of these molecules, even with varying $O_2(aq)$ tension levels.⁹⁾ Two important observations from the initial and follow-up studies were that (1) the choice of He or Ar process gas influences the resultant RONS chemistry generated by the plasma jet,⁹⁾ and (2) both He and Ar plasma jets de-oxygenate DI water.^{8,9)}

^{23, 24, 26)} De-oxygenation of the DI water with both plasma jets was attributed to the inert He or Ar gas in the partially ionized plasma jet, purging O₂(aq) out of the DI water. The importance of considering oxidation, oxygenation and de-oxygenation effects on biological cells from plasma jets was discussed in a recent study, where it was shown how each effect contributed towards the regulation of the viability of skin cells cultured *in vitro*.⁸⁾

In this study, using the UV-Vis spectroscopy, He and Ar gases and plasma jets are first compared and then it is shown how combinations of gas/plasma jet treatments can regulate the concentrations of hydrogen peroxide (H₂O₂), nitrite (NO₂⁻) and nitrate (NO₃⁻), which are the main longer-lived RONS generated by plasma in DI water,²⁷⁻³¹⁾ as well as O₂(aq). These new data are discussed in the context of how to exploit He and Ar plasma jets more effectively and efficiently in the regulation of O₂(aq) tension and RONS chemistry in aqueous solutions.

2. Experimental methods

2.1 Plasma jet

The plasma jet configuration has been described in detail elsewhere.^{9, 23-25, 32)} Briefly, the plasma jet assembly consisted of a 150 mm long, 4 mm inner diameter glass tube that was tapered to 650 μm at the nozzle. Power was supplied to a single 15 mm long external ring copper electrode wound onto the glass tube at a distance of 40 mm from the nozzle. The plasma jet assembly was used in a single electrode configuration (with no second grounded electrode). The flow rate of He or Ar through the glass tube was fixed at 0.5 standard litres per minute (slpm). During He or Ar plasma jet operation (for the free stream plasma jet) the velocity of the gas flow in both cases was 25 m/s, and the Reynolds number was calculated to be 148 and 1361, respectively (see supporting information for calculation, available online at stacks.iop.org/JJAP/58/SAAB01/mmedia). According to these Reynolds numbers, both He and Ar plasma jets were expected to operate with laminar gas flow.³³⁾ However, the Ar plasma jet appeared more filamentary, which would have generated additional gas turbulence and mixing of the plasma effluent with the ambient air. With the plasma off, the neutral gas velocity was expected to be 10-30% lower.³⁴⁾ A high voltage sinusoidal wave pulse of 10 kV_{p-p} (peak-to-peak) at 30 kHz was applied to the external electrode with a PVM500 power supply (Information Unlimited: New Hampshire, USA). A 5 mm thick

polytetrafluoroethylene (PTFE) housing was used to shield the high voltage electrode for safety. Under the parameters described above, the length of the free stream plasma jet, as seen with the unaided eye and measured with a ruler, was ≈ 7 mm and ≈ 5 mm for the He and Ar plasma jet, respectively. Voltage and current waveforms of the plasma jets were measured with a high voltage probe (Tektronix, model # P6015A) and a current monitor (Pearson, model # 2877) respectively.

2.2 Plasma-activated water

Deionized water (DI water) ($18.2 \text{ M}\Omega\cdot\text{cm}$ at $25 \text{ }^\circ\text{C}$) was obtained from a Millipore Direct-Q 3UV system. Approximately 500 mL of DI water was stored in a container in ambient atmosphere for one day before experiments. This storage procedure enabled the concentrations of dissolved gases within the DI water and the temperature to equilibrate with the air, which ensured that the concentrations of dissolved gases were constant between experiments.

2.3 UV-vis spectroscopy

A conventional double-beam UV-Vis spectrophotometer (U-3900, Hitachi) was used to measure the UV absorbance of 4.1 mL DI water inside the quartz cuvette with a standard optical path of 10 mm. The plasma jet assembly, including an x - y - z positioning stage, was mounted on the spectrophotometer. The nozzle of the plasma jet assembly was held 5 mm from the top of the DI water (Scheme and photograph in Fig. 1). The impact of the gas flow during plasma jet or neutral gas flow treatments created a slight indentation on the surface of the DI water, which increased the actual treatment distance by a negligible amount of < 1 mm. In addition, evaporation of DI water during the various treatments can further decrease the water level. The maximum amount of DI water that was evaporated for any of the experiments in this study was 13.6% - this was calculated for DI water treated with the plasma jet for 60 min, which corresponded to an approximate 4 mm decrease in the DI water level. The typical plasma jet treatment times of 15 min in this study resulted in a decrease in the water level of 1 mm. According to Henry's Law the water temperature will change the solubility of gases, which can influence the plasma jet delivery of RONS in the DI water.^{35, 36)} The influence of temperature is discussed later.

The plasma jet was aimed downward onto the DI water and real-time changes in the UV absorption profile were recorded during plasma jet treatment. The detection position was 35 mm below the water surface level (Fig. 1), which we estimate is equivalent to a 1 min delay in receiving the absorption signal (i.e. the time taken for the RONS to diffuse through the water to the detection area). As seen with the unaided eye, the He plasma jet contacted the DI water up to 10 min of treatment, but lost contact between (10-15 min); whereas the shorter length Ar plasma jet never contacted the water surface. The exact same operational parameters were also employed for the He and Ar gas treatments with the only difference being no applied voltage. The UV absorbance was used to calculate the concentrations of H_2O_2 , NO_2^- , NO_3^- and $\text{O}_2(\text{aq})$, according to an established curve-fitting routine using reference spectra of known concentrations of H_2O_2 , sodium nitrite (NaNO_2) and nitric acid (HNO_3) solutions.⁹⁾ A diagrammatical representation of the complete experimental set-up is shown in Fig. 1.

The pH and temperature of the DI water, treated under the same experimental parameters used in UV-Vis spectroscopy, was measured with a commercially available pH/temperature meter (Model D-71, Horiba).

All data points are representative of three replicate experiments ($n = 3$) and error bars are \pm standard deviations of the replicates.

3. Results and discussion

3.1 Discharge current and voltage

The applied voltage waveforms, discharge currents (I_{Dis}) and input powers (a.k.a. discharge power) for the He and Ar plasma jets are shown in Fig. 2. The sinusoidal waveforms were identical between the He and Ar plasma discharges. The peak I_{Dis} and input powers were higher for Ar compared to He. But the averaged input power over one discharge period of 33 μs was similar between He and Ar at ~ 0.47 W. Therefore, the average input energies for He and Ar, calculated after 15 min of operation, which was the usual treatment time investigated in this study, were 351 J and 354 J for He and Ar, respectively. However, the input power can vary between He and Ar plasma jets depending on the experimental parameters, as noted in a previous publication.⁹⁾ Because the input power was similar between He and Ar plasma jets in this study, the effect of input power on any differences in

RONs delivery can be discounted.

3.2 UV-vis

Employing the set-up shown in Fig. 1, UV-Vis spectroscopy was utilized to quantify the concentrations of RONS in the DI water in real-time employing different combinations of He or Ar gas and plasma jet treatments. In Fig. 3, DI water was first treated with either a He or Ar plasma jet for 15 min before the applied voltage and gas flow were extinguished. UV-Vis measurements were first taken for 15 min during the plasma jet treatments and for 45 min after the plasma and gas flow were extinguished. The results are shown in terms of the total UV absorbance (A_T , sum of absorption signal intensity in the scanning wavelength range between 190 and 340 nm) and total concentrations of H_2O_2 , NO_2^- , NO_3^- and $O_2(aq)$ in the DI water [where A_T is accounted for by these RONS and $O_2(aq)$]. For both Ar and He, it can be seen that there was an almost immediate and monotonic increase in the A_T , H_2O_2 , NO_2^- and NO_3^- , indicating fast solvation of these RONS into the DI water during plasma jet treatment. The Ar plasma jet delivered a higher concentration of H_2O_2 , whereas the He plasma jet delivered more NO_3^- , and a similar concentration of NO_2^- was delivered by both plasma jets. Both plasma jets reduced the $O_2(aq)$ concentration in the DI water to a similar level, although the recovery in $O_2(aq)$ (post plasma jet treatment) did vary. De-oxygenation of the DI water is attributed to the inert gas component of the partially ionized plasma jets, purging $O_2(aq)$ out of the DI water as previously discussed.^{8,9)} The differences observed (here) in the delivery of RONS between He and Ar plasma jets do not corroborate previously published data.⁹⁾ This discrepancy is because in terms of differences in the plasma operating parameters. When the plasma and gas flows were extinguished it can be seen that the concentration of RONS remained constant, simply because the plasma jets were no longer delivering RONS; but at the same time the $O_2(aq)$ concentration in the DI water began to increase. This increase in $O_2(aq)$ results from O_2 from the ambient air, solvating back into the DI water.

The ability of the plasma jets to deliver stable RONS (from the reactive gas component) and simultaneously to deoxygenate (from the inert gas component) could potentially be exploited to modulate the ratio of RONS/ $O_2(aq)$ in the DI water. This idea was directly tested for both He and Ar in the next experiment, where DI water was first treated for 15 min with

either an Ar or He plasma jet before the applied voltage (i.e. plasma) was extinguished and the (respective) inert gas jet was applied for a further 45 min. The results are shown in Fig. 4. The trends in the delivery of RONS and de-oxygenation are similar to those described from Fig. 3 for the first 15 min of plasma jet treatment. When the plasma jets were extinguished and gas jets left on, the concentrations of RONS once again stabilized in the DI water because the plasma jets were no longer delivering RONS (consistent with Fig. 3). However, in both cases the $O_2(aq)$ concentration continued to decrease and at rates that exceeded those for when the Ar and He plasma jets were ignited as expected from the literature.^{8,9,37)} Nearly all of the $O_2(aq)$ was removed from the DI water at the ~ 30 min time-point by the He gas jet. At the ~ 25 min time-point the Ar gas jet had reduced the $O_2(aq)$ concentration to $\sim 2\text{-}3\text{ mgL}^{-1}$; this was the lowest $O_2(aq)$ concentration that could be achieved with the Ar gas jet and pro-longed treatment did not further reduce the $O_2(aq)$. Occasionally, the RONS concentrations continued to increase during inert gas treatments for DI water initially treated with the plasma jets (e.g. see H_2O_2 graph in Fig. 4).

It should be noted that the RONS and $O_2(aq)$ concentrations did not always precisely match between experiments. This is presumably due to experimental variation or small errors in the automated curve-fitting routine of the UV-Vis data as noted in a previous publication.⁹⁾ These errors might also arise from the highly complex physicochemical processes occurring particularly near the surface of the DI water during plasma jet treatments, which lead to generation of further RONS not accounted for in this study³⁸⁾ – in this study we focus only on the major longer-lived RONS generated by atmospheric plasma in water. In addition, post-plasma jet treatment we observed that the H_2O_2 concentration generally continued to slightly increase (see Figs. 3-5), whereas the concentrations of NO_2^- or NO_3^- remained more-or-less constant. These results are different to the results obtained by Lukes *et al* who observed a decrease in H_2O_2 and NO_2^- , but an increase in NO_3^- post plasma treatment.³⁹⁾ This could be due to the very different plasma sources utilized by Lukes *et al* (air discharge plasma) compared to the He/Ar plasma jets in this study. Based on the study by Lukes *et al*³⁹⁾ we note that H_2O_2 can also be formed through reactions between hydroxyl radicals also generated by the plasma jets. Also, we appreciate that changes in pH will influence the production of RONS post-plasma jet treatment.⁴⁰⁾ However, further detailed experiments will need to be performed to develop a more comprehensive understanding of the liquid

chemistry. In the future the authors will address this issue; but for the present study the general trends in RONS/ $O_2(aq)$ delivery were more important than absolute values.

The role played by gas flow is made clearer when in Fig. 5 UV-Vis was used to monitor RONS and $O_2(aq)$ in DI water for 60 min with plasma jet treatments for the first 15 min, followed immediately after with 15 min inert gas jet treatments, and the 30 min with no applied voltage and gas flow. The trends in RONS delivery and de-oxygenation are identical to those seen in Fig. 3 and Fig. 4, with a notable exception being that the trend for NO_3^- delivery was similar for He and Ar plasma jets, whereas in Fig. 3 and Fig. 4 the He plasma jet delivered more NO_3^- compared to the Ar plasma jet. As mentioned before, this discrepancy is due to experimental variation and/or small errors in the automated curve-fitting routine. Immediately after plasma jet treatment, between time-points 15-30 min, it can be seen that the He gas jet was more effective at de-oxygenating the DI water compared to the Ar gas jet (consistent with Fig. 4). Following treatment with the inert gas jets, and with the plasma and gas flow both off, the $O_2(aq)$ concentration began to recover in the DI water; and at the 60 min time-point the $O_2(aq)$ concentration was the same for He and Ar at $\sim 4 \text{ mg L}^{-1}$ but still approximately 50% lower compared to the original $O_2(aq)$ concentration.

So far, the inert gas treatments of DI water were performed immediately after the plasma jet treatments. Therefore, the effect on the RONS and $O_2(aq)$ concentrations of the gas flows only was next investigated. In Fig. 6 He and Ar gas jets were applied for 30 min, and UV-Vis was used to follow $O_2(aq)$ recovery for the next 30 min without gas flow. As expected, in this control experiment the inert gas jets alone do not change the RONS concentrations in the DI water (consistent with Fig. 4 and Fig. 5). Again, as expected, the extent of de-oxygenation seen with the He gas jet was greater than the Ar gas jet. When the inert gas jets were switched off, the $O_2(aq)$ concentration immediately began to recover (consistent with Fig. 5). It can also be seen that the decrease and increase in the A_T follows a similar trends for the decrease and increase in $O_2(aq)$. As already discussed above, this is expected because $A_T = \text{total RONS} + O_2(aq)$ concentrations in DI water. The decrease in A_T corresponding to the decrease in $O_2(aq)$ was not revealed in other graphs (e.g. in Fig. 4 and Fig. 5) because of the larger intensity units required to plot A_T when RONS are generated; this is because of the lower co-efficient of absorbance of $O_2(aq)$, and consequently A_T is less sensitive to changes in $O_2(aq)$ when H_2O_2 , NO_2^- and NO_3^- are present.⁹⁾

Reversing the order of the previous other plasma jet experiments, Ar and He gas jets were used to first de-oxygenate the DI water followed immediately by plasma jet treatments. The treatment times were: 15 min inert gas jet, 15 min plasma jet, 30 min plasma jet and gas flow off. The results are shown in Fig. 7. Although the inert gas jet reduced the original O₂(aq) concentration to ~ 25% and ~ 50% (of original levels) in DI water for He and Ar, respectively, the low O₂(aq) concentration did not affect the generation of RONS with the plasma jets; similar trends and concentrations were observed to those in Figs. 3-5 in the de-oxygenated DI water. But in contrast to oxygenated DI water equilibrated with ambient air, with de-oxygenated DI water, both plasma jets increased the concentration of O₂(aq). The reason for the increase in O₂(aq) could possibly be due to the decay of RONS giving rise to O₂(aq) such as $2\text{HOO}\cdot \rightarrow \text{H}_2\text{O}_2 + \text{O}_2(\text{aq})$ or O₂ from the ambient air being entrained in the effluent of the plasma jet gas and consequently solvating into the DI water.

Having analyzed changes in the concentrations of RONS and O₂(aq) induced by He and Ar gas and plasma jets, the next experiments were designed to test if combinations of these treatments can be used to precisely regulate the oxygen tension and RONS concentrations in DI water. This could be useful if gas and plasma jets are utilized to regulate cell growth or function in bioreactors; e.g. decreasing the oxygen tension should decrease cell metabolism and reduce cell growth, whereas small increases in RONS concentrations can enhance cell proliferation.^{8,41)} For these experiments, the treatments were as follows: 9 min inert gas jet, 9 min plasma jet, 9 min inert gas and plasma jet off. The series of treatments was repeated twice and UV-Vis of the results were recorded for 60 min. These results are shown in Fig. 8. It can be seen that the concentration of RONS and O₂(aq) was controlled with relatively good precision. Successive treatments with the inert gas jets resulted in well-defined decreases in O₂(aq) of similar magnitude, whereas successive plasma jet treatments resulted in well-defined incremental increases in the RONS concentrations.

In the scale-up to industrial applications it is also important to consider not just the control but also the efficiency of the plasma jets in the production of RONS in liquid. A protocol was previously established for assessing the efficiency of plasma jet generation of RONS in DI water.²²⁾ In Ref. 22 it was shown that plasma jets operated with a micron-sized nozzle are more efficient at RONS generation compared to millimeter-sized nozzles, under otherwise identical conditions of operation. A second important observation in Ref. 22 was

that subtle changes in the treatment distance over millimeter scale-lengths between the nozzle and target solution, significantly impacted on the efficiency of RONS generation; but the relationship was not linear. Herein, the same protocol described in Ref. 22 was employed in this study to compare the efficiency of RONS generation for He and Ar plasma jets with a micron-sized nozzle of 650 μm . These results are presented in Fig. 9 with production efficiency expressed as $\text{mg L}^{-1} / \text{J}$: where mg = quantity of RONS; L = volume of water; and J = input energy. Displaying the data in this manner it is clear that the Ar plasma jet was more efficient at producing H_2O_2 , the He marginally more efficient at producing NO_3^- and both equally efficient at producing NO_2^- . Because the efficiency in the plasma jet delivery of RONS was shown to vary for different RONS and between He and Ar process gases and also the different treatment condition; e.g. He plasma jet is in contact with the water surface while Ar plasma jet is not contact with the liquid. The data show it is important to consider the production efficiency of RONS with plasma jets for larger scale industrial applications, where efficiency of RONS generation is important for maintaining lower production costs.

One aspect this study touches upon is the need to improve reliability (or predictability) in the modulation of RONS chemistry that is achieved with different plasma jets or other plasma sources, according to mode of operation or plasma set-up; e.g. the discrepancies between the results in this study to Lukes *et al.*³⁹⁾ Further discrepancy can be seen when comparing results herein with one of our own previous studies⁹⁾ in focusing on the plasma jet delivery particularly of NO_2^- and NO_3^- . In order to highlight this discrepancy, a comparison is made between the He and Ar plasma jet delivery of RONS in this study (Fig. 3 above) to Fig. 9 in Ref. 9. Both studies employed a similar plasma jet set-up; in both studies 4.1 mL of DI water in a quartz cuvette was treated for 15 min. Major physical (known) differences are the nozzle inner diameter, the applied voltage and frequency, treatment distances and whether or not the plasma jet plume contacted the surface of the DI water. In Ref. 9 both He and Ar plasma jets never contacted the surface of the DI water. Changing these parameters resulted in striking differences in RONS delivery between He and Ar plasma jets. Although other differences in experimental parameters cannot be discounted, we think a major contributing factor to these discrepancies is the treatment distance. This is based upon the study of Wende *et al* who observed that treatment distance affected the delivery of RONS from one particular plasma source but not for a second but similar plasma

source.⁴²⁾ Currently, we can only speculate on why the resultant RONS chemistry in the DI water is particularly sensitive to the treatment distance. Different treatment distances will result in different amounts of the plasma effluent mixing with air (before the gas flow reaches the liquid surface), and treatment distance will also affect water evaporation. For an example, we showed that the RONS chemistry is tunable by simply changing the treatment distance in previous work.⁴³⁾ Both of these points will significantly influence the production of RONS in the gas phase, which will affect the resultant liquid chemistry. In addition, variabilities in the voltage, frequency and nozzle diameter may also influence the RONS chemistry and should not be discounted; in particular these differences can influence length of the plasma plume, which could have significant effect on RONS production (refer to discussion in following paragraph). Further fundamental research is required to establish how the plasma jet experimental parameters influences the RONS chemistry. This is not trivial especially owing to the complexity of nitrogen-oxide liquid chemistry, which is only partially understood.^{31,44)}

To provide further insight into the mechanisms of plasma-induced RONS chemistry in solution, the temperature and pH of the water was measured at different time-points of He and Ar plasma jet treatments (Fig. 10). In Fig 10(a), it can be seen that the pH of the DI water decreased for both plasma jets with a greater decrease seen for the He plasma jet. Forty five minutes after treatments, the pH of the solutions increased again as the DI water began to equilibrate with the ambient air. In Fig. 10(b) it is seen that the water temperature increases with He plasma jet treatment times up to 10 min, but begins to decrease after pro-longed treatment of 15 min. For the Ar plasma jet, the water temperature remained relatively constant, and if anything slightly decreased, for all treatment times tested [Fig. 10(b)]. Forty five minutes after plasma treatments, the water temperature returned to approximately the original value. We propose that these differences in behavior between He and Ar plasma jets are attributed to the differences in the length of the plasma jets; ≈ 7 mm and ≈ 5 mm for the He and Ar plasma jet, respectively. For these plasma jets, the He plasma jet contacted the water surface for up to the first 10 min of treatment. But by 15 min, plasma jet contact with the water surface was lost. This was due to the water evaporation lowering the water level in the cuvette. However, the shorter Ar plasma jet never contacted the water surface. Therefore, in the case of He, in contact mode, we propose that electrons and ions emanating

from the plasma jet directly heat the water for up to 10 min. At 15 min, the water temperature decreases due to evaporative cooling. Since the Ar plasma jet was never in contact with the water surface, the water temperature remained relatively constant with a slight decrease (if anything) due to evaporation. Based upon on the literature we propose that the higher temperature seen for the He plasma jet enhances NO_x reactions leading to formation of nitric acid and a lowering of the water pH.⁴⁰⁾

A number of different molecules are detected in water treated by plasma jets including short and longer lived molecules. We have established our UV-Vis method to accurately detect H_2O_2 , NO_2^- , NO_3^- and $\text{O}_2(\text{aq})$. We do not discount the possibility that other molecules may also contribute to the UV absorbance. For example, it is important to understand the chemistry responsible for lowering the pH of the water during plasma jet treatment, as well the chemistry during the increase (recovery) of the pH after the plasma is switched off. The lowering of the pH solution during plasma jet treatment is likely to be due to formation of nitric acid, which we did not account for in the deconvolution of the UV absorbance spectra. However, the main decomposition product from nitric acid is NO_3^- ,⁴⁵⁾ which we did measure in this study. To accurately quantify further molecules by our UV-Vis will require analysis of the reference spectrum at different wavelengths for each individual molecule at different concentrations, as detailed in a previous publication.⁴³⁾

In summary, this study not only demonstrated the efficacy but also highlighted the remaining challenges in using He and Ar plasma jets to modulate the $\text{O}_2(\text{aq})$ tension and RONS chemistry in an aqueous target. It is important to address these challenges because the development of He and Ar plasma jets to modulate RONS and oxygen tension in aqueous solutions with a high level of reproducibility, reliability and accuracy, could be exploited in a number of industrial and medical applications. This includes nanomaterials synthesis,⁴⁶⁾ chemical processing,⁴⁷⁾ regulation of cell growth in bioreactors for biotechnological and pharmaceutical industries and in the production of cells for cell therapy.^{41, 48-51)}

4. Conclusions

Different combinations of He and Ar plasma and gas jet treatments were used to modulate the $\text{O}_2(\text{aq})$ tension and RONS chemistry in DI water. He and Ar plasma jets delivered oxidizing species (H_2O_2 , NO_2^- and NO_3^-) and oxygenated or de-oxygenated DI water

depending on the $O_2(aq)$ concentration at the time of treatment. He and Ar gas jets deoxygenated the DI water. Considering the electrical energy consumption, the Ar plasma jet more efficient at producing H_2O_2 , and both equally efficient at producing NO_2^- . Improving our knowledge of the nitrogen-oxide liquid chemistry should assist the development of accurate, reproducible and efficient plasma jet methods for modulating RONS and $O_2(aq)$ in aqueous solutions. This could significantly benefit industries that utilize large-scale chemical processing facilities or cell culture bioreactors for the manufacture of pharmaceutical products and cells for cell therapy.

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Figures and Figure Captions

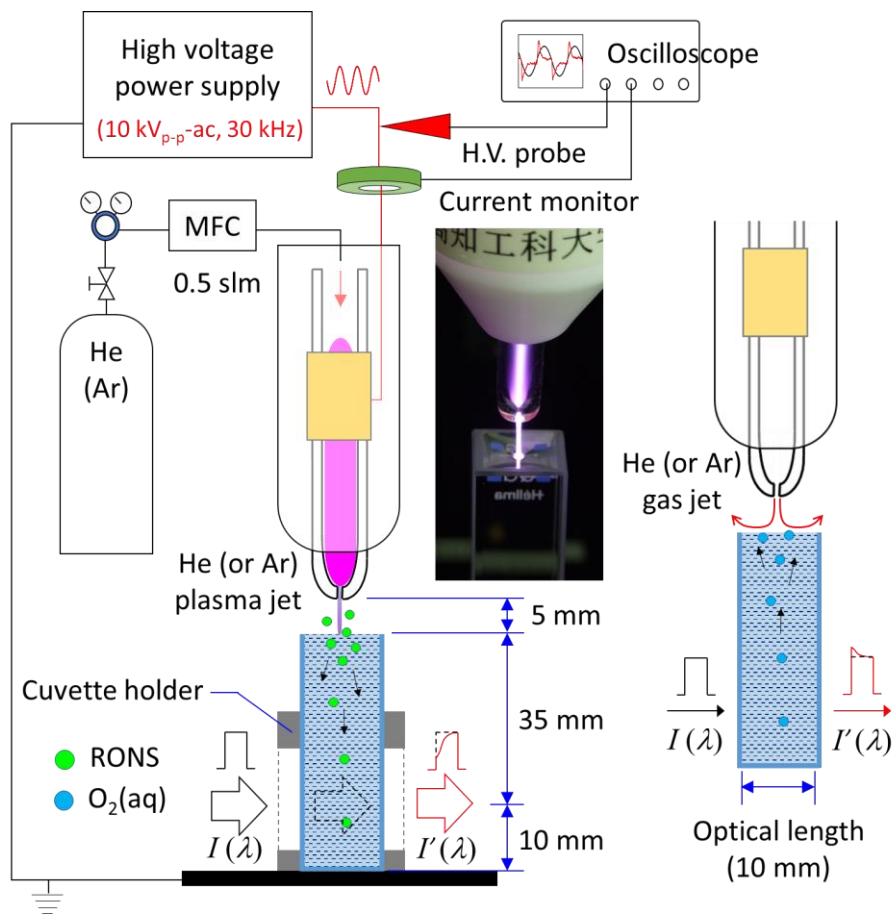


Fig. 1. (Color online) Experimental set-up where UV-Vis spectroscopy is used to monitor real-time changes in RONS and $O_2(aq)$ during He and Ar plasma and gas jet treatment of DI water.

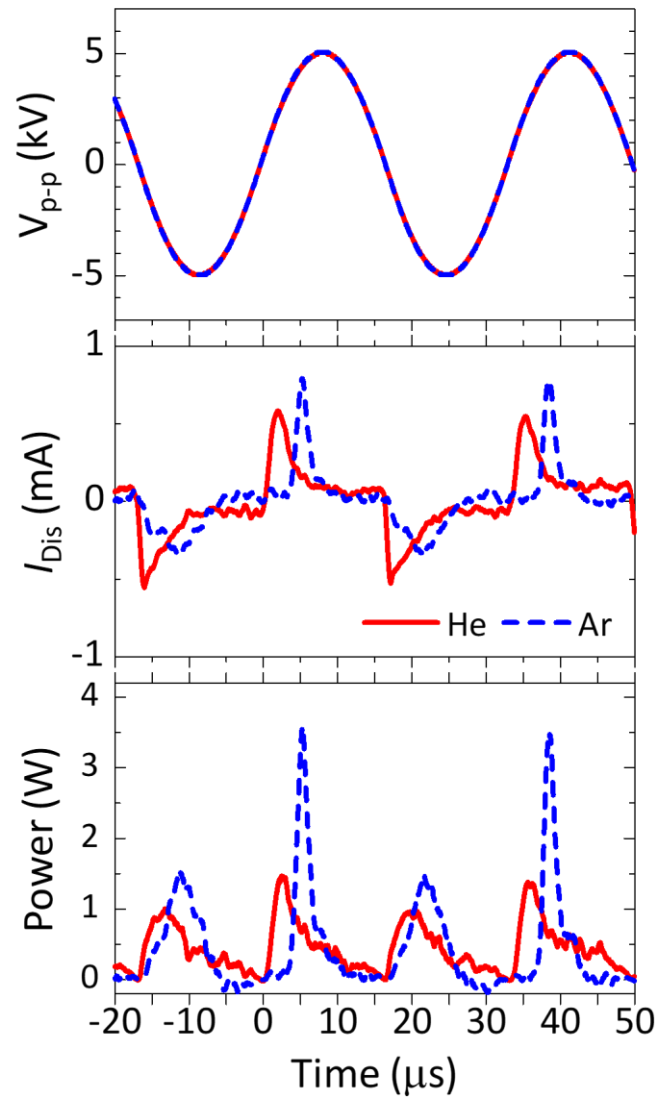


Fig. 2. (Color online) Electrical characteristics of the He and Ar plasma jets. From top-to-bottom, the plots show (top graph) the peak-to-peak applied voltage waveform at 10 kV_{p-p} with a frequency of 30 kHz, (middle graph) discharge current (I_{Dis}), and (bottom graph) the input power of the discharge. The red solid lines and blue dashed lines are for the He and Ar plasma jets, respectively (as indicated in legend on middle graph).

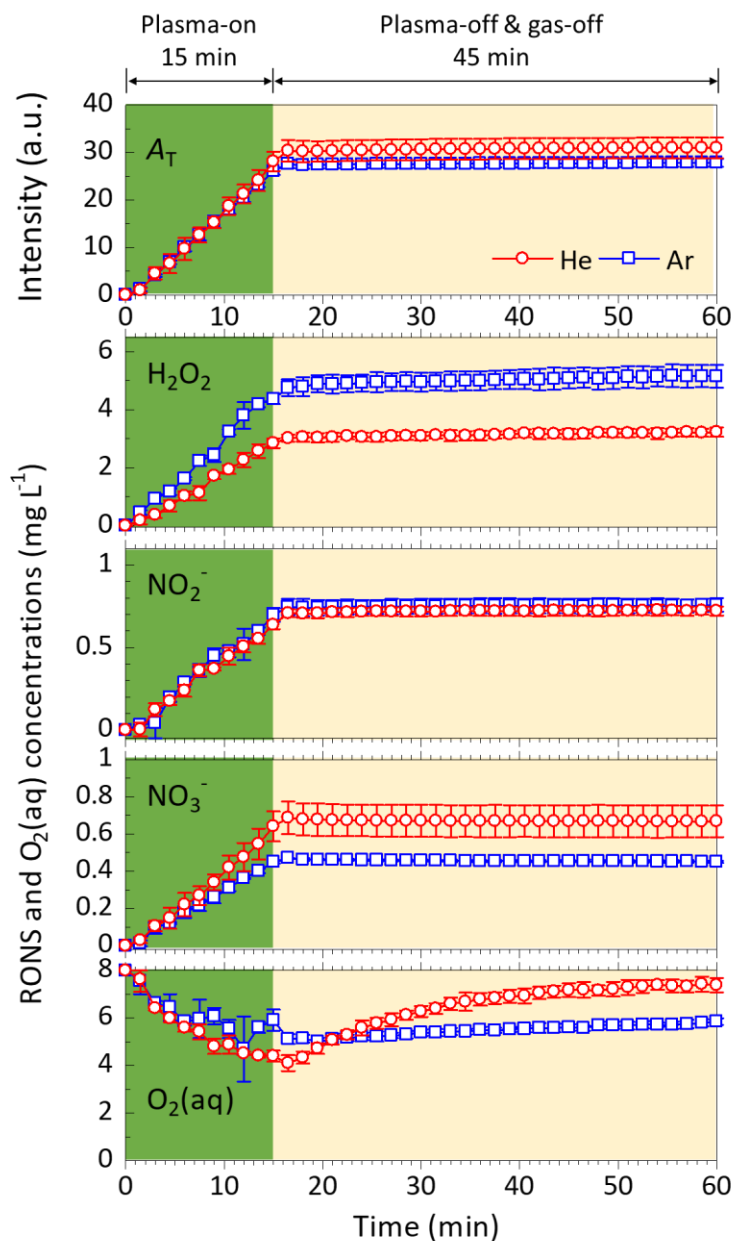


Fig. 3. (Color online) Time-dependent change in the A_T and concentrations of H_2O_2 , NO_2^- , NO_3^- , and $O_2(aq)$ in DI water during He and Ar plasma jet treatments, as measured by UV-Vis. The treatment time was 15 min (shaded green) after which the plasma jets and the gas flows were switched off (shaded beige). The red circles and blue squares represent the data for the He and Ar plasma jets, respectively (as indicated in legend on top graph). A_T = sum of absorption signal intensity in the scanning wavelength range between 190-340 nm.

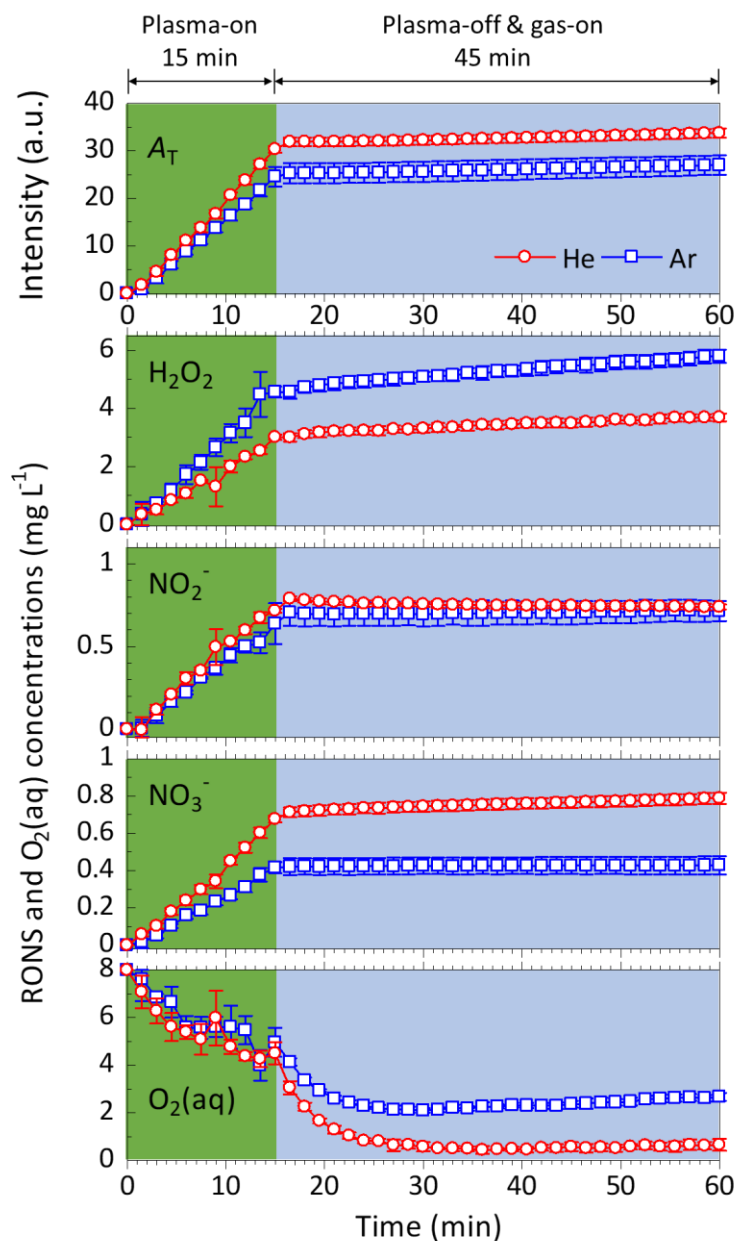


Fig. 4. (Color online) Time-dependent change in the A_T and concentrations of H_2O_2 , NO_2^- , NO_3^- , and $O_2(aq)$ in DI water during He and Ar plasma jet and gas jet treatments, as measured by UV-Vis. The plasma jet treatment time was 15 min (shaded green), followed immediately by 45 min of inert gas jet treatment (shaded blue). The red circles and blue squares represent the data for the He and Ar plasma or gas jets, respectively (as indicated in legend on top graph). A_T = sum of absorption signal intensity in the scanning wavelength range between 190-340 nm.

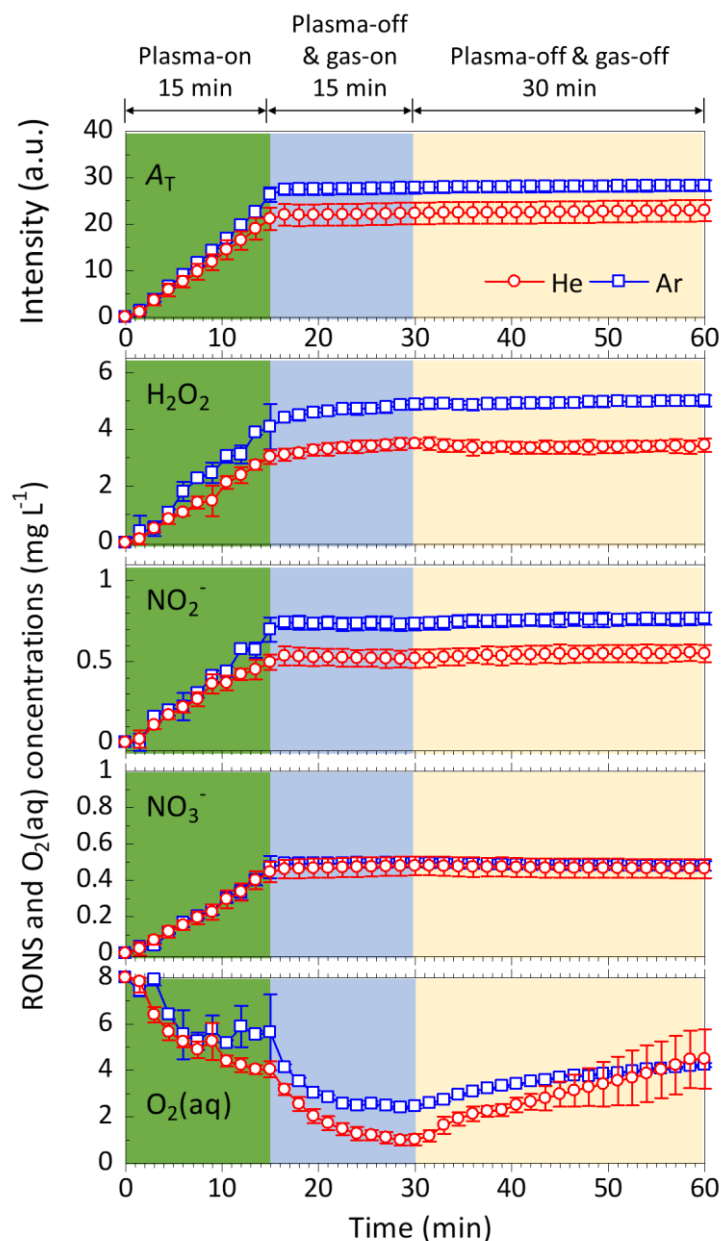


Fig. 5. (Color online) Time-dependent change in the A_T and concentrations of H_2O_2 , NO_2^- , NO_3^- , and $O_2(aq)$ in DI water during He and Ar plasma jet and gas jet treatments, as measured by UV-Vis. The plasma jet treatment time was 15 min (shaded green), followed immediately after with 15 min of inert gas jet treatment (shaded blue), and 30 min with the plasma jet and gas flow off (shaded beige). The red circles and blue squares represent the data for the He and Ar plasma or gas jets, respectively (as indicated in legend on top graph). A_T = sum of absorption signal intensity in the scanning wavelength range between 190-340 nm.

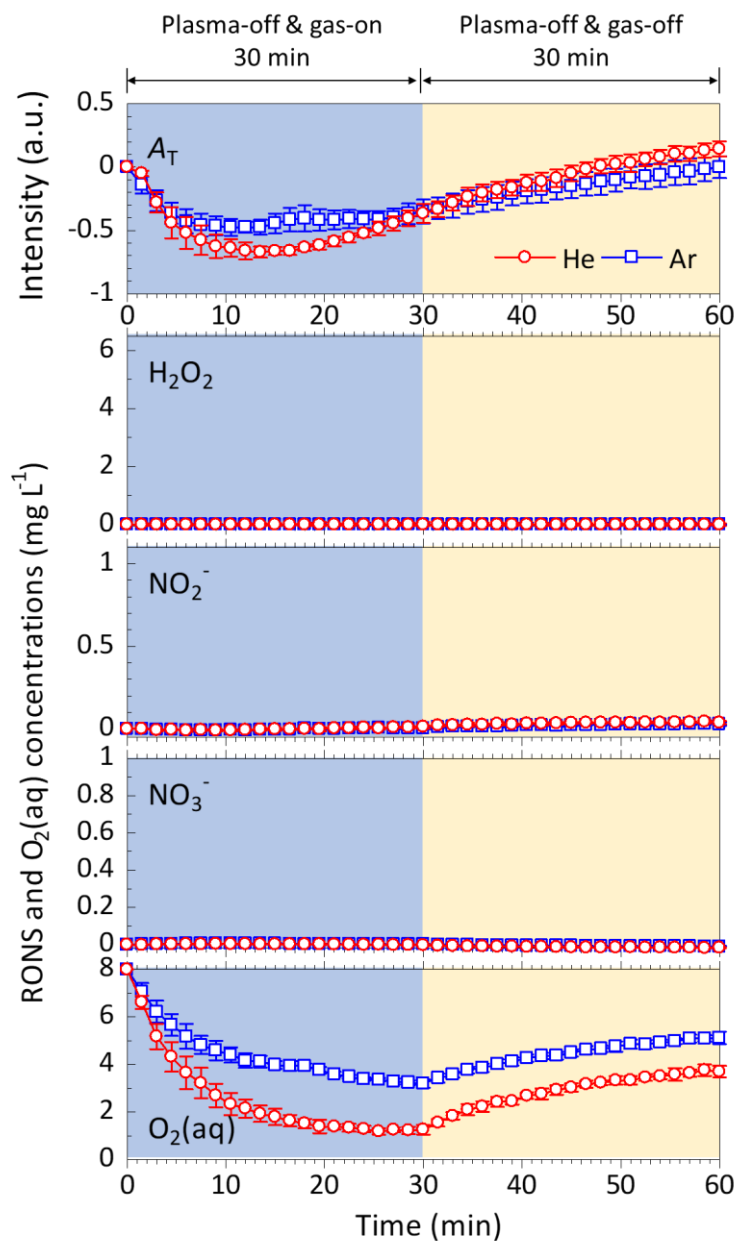


Fig. 6. (Color online) Time-dependent change in the A_T and concentrations of H_2O_2 , NO_2^- , NO_3^- , and $O_2(aq)$ in DI water during He and Ar gas jet, as measured by UV-Vis. The treatment time was 30 min (shaded blue), and 30 min with the gas flow off (shaded beige). The red circles and blue squares represent the data for the He and Ar gas jets, respectively (as indicated in legend on top graph). A_T = sum of absorption signal intensity in the scanning wavelength range between 190-340 nm.

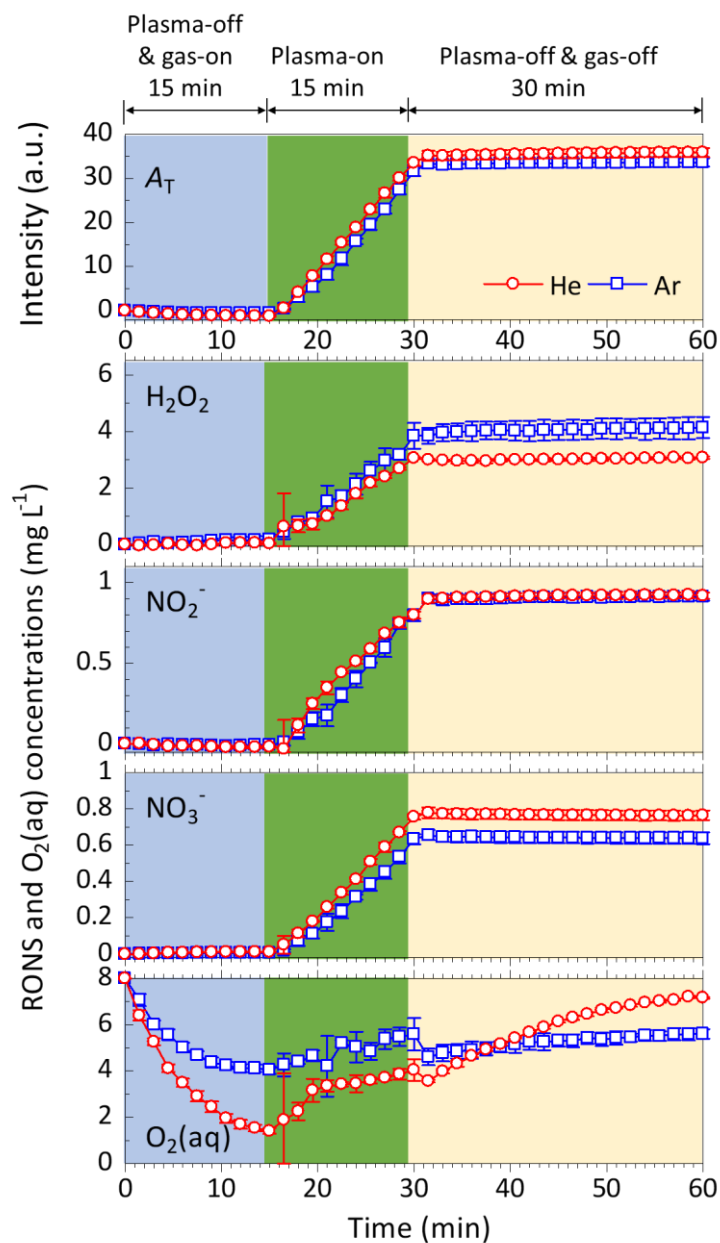


Fig. 7. (Color online) Time-dependent change in the A_T and concentrations of H_2O_2 , NO_2^- , NO_3^- , and $O_2(aq)$ in DI water during He and Ar gas jet and plasma jet treatments, as measured by UV-Vis. The gas jet treatment time was 15 min (shaded blue), followed immediately after with 15 min of plasma jet treatment (shaded green), and 30 min with the plasma jet and gas flow off (shaded beige). The red circles and blue squares represent the data for the He and Ar gas and plasma jets, respectively (as indicated in legend on top graph). A_T = sum of absorption signal intensity in the scanning wavelength range between 190-340 nm.

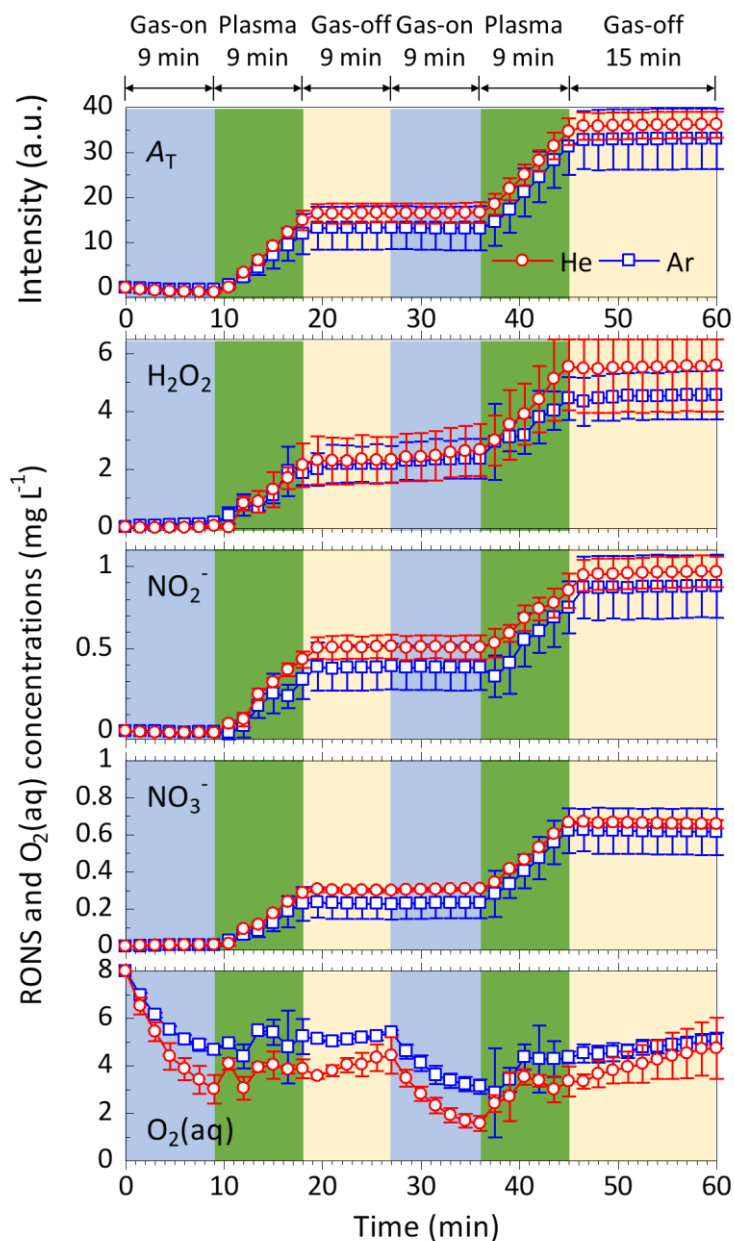


Fig. 8. (Color online) Time-dependent change in the A_T and concentrations of H_2O_2 , NO_2^- , NO_3^- , and $O_2(aq)$ in DI water during repetitive He and Ar gas jet and plasma jet treatments, as measured by UV-Vis. The gas jet treatment time was 9 min (shaded blue), followed immediately after with 9 min of plasma jet treatment (shaded green), and 9 min with the plasma jet and gas flow off (shaded beige). The treatment cycle was repeated twice with UV-Vis recorded to the 60 min time-point. The red circles and blue squares represent the data for the He and Ar gas and plasma jets, respectively (as indicated in legend on top graph). A_T = sum of absorption signal intensity in the scanning wavelength range between 190-340 nm.

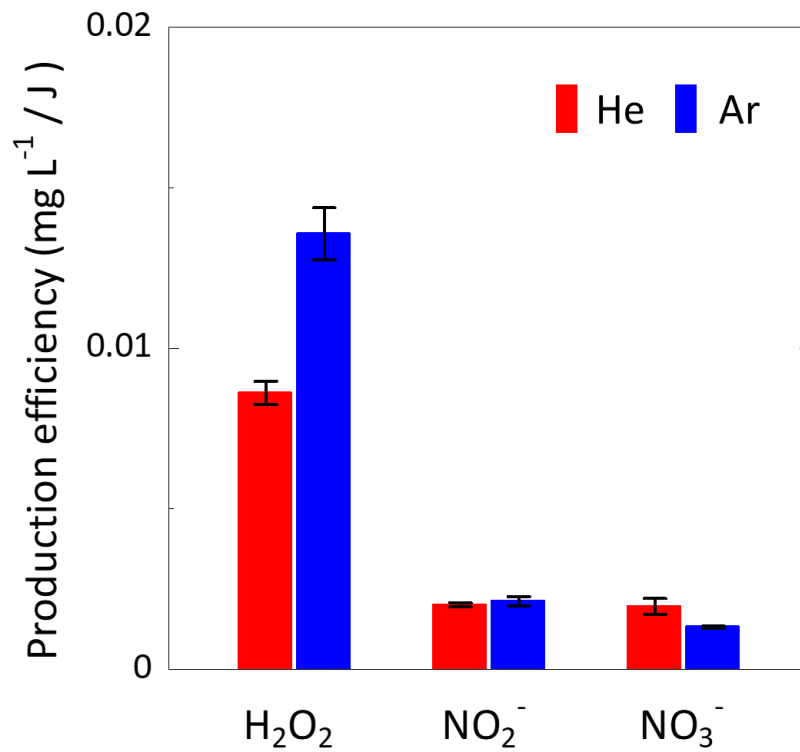


Fig. 9. (Color online) Production efficiency of He and Ar plasma jets for the delivery of H₂O₂, NO₂⁻, and NO₃⁻ into DI water. Efficiency of RONS production was measured after 15 min plasma jet treatment of 4.1 mL DI water as per the experimental set-up in Fig. 1. The red and blue bars represent the data for the He and Ar plasma jets, respectively (as indicated in the legend).

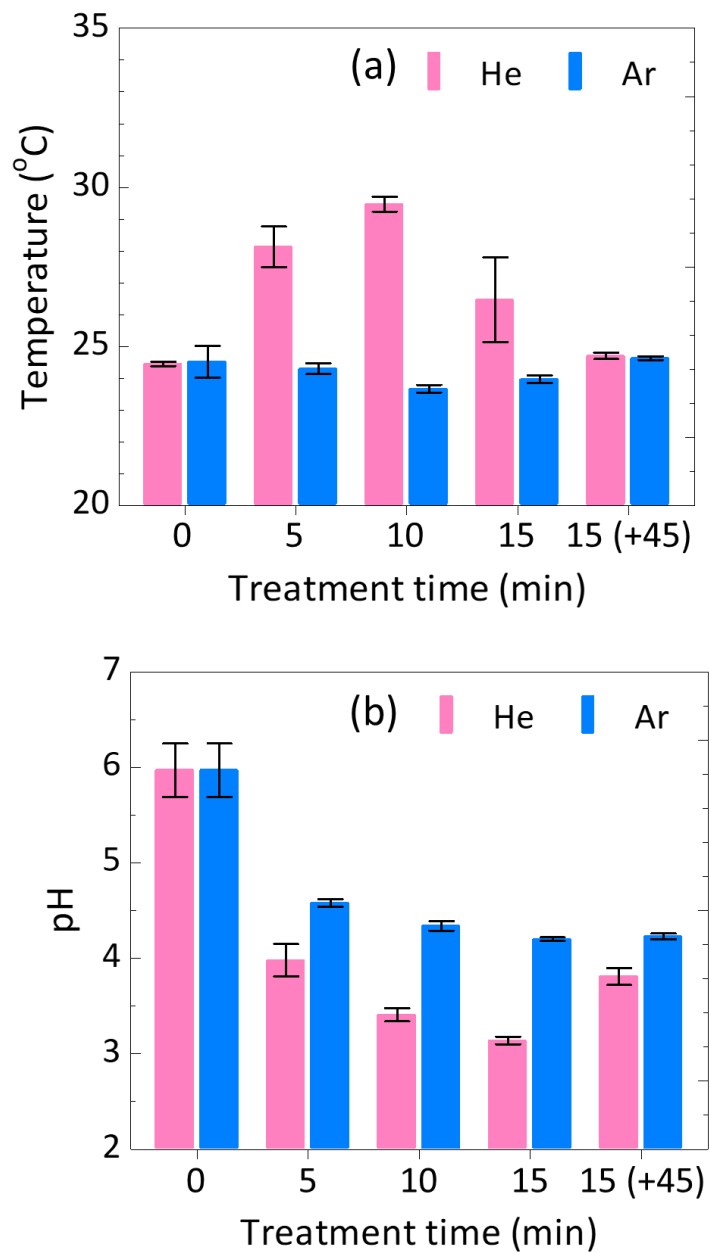


Fig. 10. (Color online) Time-dependent change in the (a) pH and (b) temperature of the DI water after different times of He and Ar plasma jet treatments. On the far right side, 15(+45) refers to measurements taken 45 min after a 15 min plasma jet treatment. The He plasma jet contacted the water up to 10 min but was not in contact at 15 min. The Ar plasma jet never contacted the water surface.