

架橋化多孔性タンパク質結晶を利用した水素製造光触媒システムの構築

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Photocatalytic hydrogen evolution systems constructed in cross-linked porous protein crystals

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<p>ポイント</p>	<ul style="list-style-type: none"> ・安価で入手しやすい鶏卵由来のリゾチーム結晶と、無尽蔵の光エネルギーを使って水素が製造できる光触媒システムを構築 ・高効率で水素を製造(水素の収率 76%→85%) ・リゾチーム結晶の内部間隙に機能性分子とナノ粒子を集積化することで、合理的な触媒システムの設計が可能に
<p>概要</p>	<p>ナノサイズの細孔をもつタンパク質(多孔性タンパク質結晶)の内部に、太陽光エネルギーを吸収する光増感剤分子と水素発生反応に対し触媒活性をもつ金属微粒子を近接させ集積化することで、光エネルギーを利用して水素製造ができるシステムの構築に成功しました。</p> <p>この成果を利用することで、従来の光触媒系に比べて高い効率での光エネルギーの有効利用が可能になると期待されます。</p>
<p>本研究成果の意義</p>	<p>本研究では、天然に存在するタンパク質を有機多孔性材料として用いて、機能性分子とナノ粒子を同時に集積化し、水素製造光触媒システムを高効率化する技術を開発しました。</p> <p>また、原子レベルでの構造観察の結果から、高効率化した理由も突き止めることができました。</p> <p>今後、それぞれの化学反応の素反応過程に適した機能性分子や材料を任意の位置に配置できるタンパク質結晶を担体として利用することで、さまざまな用途に応じた触媒システムを合理的に設計できるようになり、持続発展可能な社会の実現に繋がることが期待されます。</p>
<p>共同研究資金等</p>	<p>本研究は東京工業大学との共同研究で、下記の資金援助を得て実施されました。</p> <p>公益信託 ENEOS 水素基金「メソ結晶内部の隣接間隙を利用した複合型光水素発生触媒」(研究代表者：山田 裕介)</p>

Photocatalytic hydrogen evolution systems constructed in cross-linked porous protein crystals

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Abstract

Cross-linked hen egg white lysozyme crystals (CL-HEWL) have been employed as supports to construct heterogeneous catalysts for photocatalytic hydrogen (H₂) evolution, where rose bengal (RB) and Pt nanoparticles (PtNPs) acted as a photosensitizer and H₂-evolution catalysts, respectively. Single-crystal X-ray structure analyses of the CL-HEWL immobilizing a precursor for PtNPs suggested that a coordination site of the precursor locates in immediate proximity to potential adsorption sites for RB. The accumulation of the components facilitated photo-induced electron transfer, resulting in efficient H₂ evolution. These results suggest that porous protein crystals are promising platforms to periodically and rationally accumulate catalytic components by using molecular interactions.

Keywords: Heterogeneous catalysis; Photocatalysis; Hydrogen evolution; Protein crystal; Lysozyme

1. Introduction

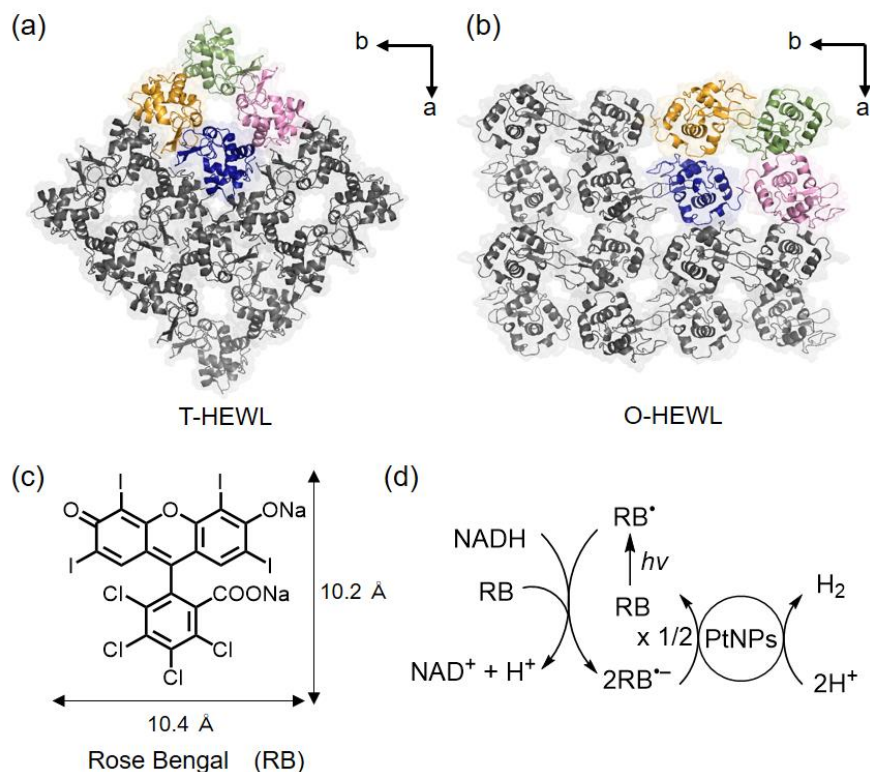
Bio-supramolecular structures utilizing proteins have attracted much attention to develop highly sophisticated functional materials [1–10], because multiple functional molecules can be immobilized simultaneously in a protein crystal by utilizing specific interaction with the functional groups of amino acids [11,12]. Additionally, interaction of side chains among proteins often induces self-assembled nanostructures available as nano-vessels for immobilization of metal clusters composed of multiple atoms [13–16]. Moreover, metal nanoparticles larger than the metal clusters can be immobilized in protein crystals where larger spaces are available among protein molecules aligned periodically [17–19]. Recently, protein crystals with a cross-link treatment have emerged as a robust porous material [20]. The pores can be provided by removal of solvent molecules from, so-called, solvent channels, in which metal complexes and organic compounds can be immobilized by selective interactions such as coordination, electrostatic interaction, hydrophobic interaction, etc, with the functional groups at the pore walls [21–24]. In fact, site-selective immobilization of metal nanoparticles by reduction of pre-organized metal complexes in solvent channels has been reported [14], indicating that sophisticated functional materials can be achieved by cooperative immobilization of functional molecules and metal nanoparticles in porous protein crystals.

Cooperative immobilization of a photoresponsible molecule and metal nanoparticles can provide solar energy conversion systems to produce high-energy compounds such as hydrogen (H_2) [25–28]. Efficient photocatalytic H_2 -evolution systems can be achieved by combination of several catalytic components including light-harvesting and charge-separation units and water-oxidation and reduction catalysts developed separately [29–38]. These components are usually combined in an appropriate solvent, however, back electron transfer between oxidation and

reduction catalysts is hardly avoidable. Immobilization of these components with arrangement suitable for forward electron transfer can enhance the photocatalytic efficiency on the surfaces of a solid support such as mesoporous silica-alumina [34–37]. However, amorphous nature of mesoporous silica-alumina disturbs precise and periodical arrangement of components, which is necessary for efficient photocatalytic H₂ evolution. In addition, anionic components are hardly immobilized on silica-alumina due to the electrostatic repulsion.

We report herein photocatalytic H₂-evolution systems constructed in cross-linked hen egg white lysozyme crystals by immobilizing Pt nanoparticles (PtNPs) as H₂-evolution catalysts in immediate proximity to an anionic photosensitizer molecule, rose bengal (RB). β -Dihydronicotinamide adenine dinucleotide (NADH) was employed as an electron donor, because irreversibly decomposing nature of NADH during the reaction can result in quantitative H₂ evolution, allowing to evaluate the efficiency of the reaction systems by H₂ yields [39]. Construction of photocatalytic H₂ evolution systems composed of molecular H₂-evolution catalyst and molecular photosensitizers in a viral protein particle has been reported [7], however, the small interior space allowed no co-immobilization of metal nanoparticles and photosensitizers. Larger inner spaces of porous protein crystals are necessary for co-immobilization of an organic photosensitizer and metal nanoparticles. Additionally, porous protein crystals immobilizing exogeneous molecules are stable in a photoreaction under visible light irradiation [40]. Thus, cross-linked hen egg white lysozyme crystals in the tetragonal form (CL-T-HEWL) and in the orthorhombic form (CL-O-HEWL) were employed as porous supports (Schemes 1a and 1b). The prepared catalysts were characterized by single-crystal X-ray structure analyses and TEM observations. Photocatalytic H₂ evolution was performed in the presence of NADH under visible-

light irradiation of a suspension containing the composite photocatalysts. The chemical structure of RB and the overall photocatalytic cycle



Scheme 1. Porous structure of HEWL crystals in (a) tetragonal form (T-HEWL) and (b) orthorhombic form (O-HEWL) from the Protein Data Bank (PDB), codes: 193L and 1BGI, respectively. (c) Chemical structure of rose bengal (RB) and (d) the overall photocatalytic cycle of H_2 evolution.

are indicated in Schemes 1c and d, respectively. The accumulation of the components resulted in increasing the amount of evolved H_2 to three times that of a homogeneous system without the HEWL crystals.

2. Experimental section

2.1. Materials

Hen egg white lysozyme (HEWL) was purchased from Sigma–Aldrich Co. LLC. Other reagents were purchased from Tokyo chemical industry Co., Ltd., Wako pure chemical, and Sigma–Aldrich and were used without further purification.

2.2. Catalyst characterization

UV-vis absorption spectra of rose bengal (RB) on each CL-T-HEWL composite were measured using a JASCO V-770 with an integrating sphere attachment. The atomic ratio of sulfur, iodine and platinum in each CL-T-HEWL composite was determined using a Shimadzu EDX–730 X-ray fluorescence spectrometer. Transmission electron microscope (TEM) images of PtNP@RB·CL-T-HEWL were obtained using a JEOL JEM–2100 equipped with a field-emission gun with an accelerating voltage of 200 kV. Thin pieces of PtNP@RB·CL-T-HEWL were fixed on a Cu-mesh microgrid coated with an amorphous carbon supporting film.

2.2.1. Single-crystal X-ray structure analysis

The hanging drop vapor diffusion method was used to obtain HEWL crystals in the tetragonal form and in the orthorhombic form followed by the cross-linking treatment (CL-T-HEWL and CL-O-HEWL, respectively) [41,42]. These crystals were soaked in a drop of HEPES buffer (10 mM, pH 7.8) containing RB (10 mM) and NaCl (1.0 M) at room temperature for 24 h (RB·CL-T-HEWL and RB·CL-O-HEWL, respectively) followed by soaking in a drop of HEPES buffer (10 mM, pH 7.8) containing $\text{H}_2\text{PtCl}_6 \cdot 6\text{H}_2\text{O}$ (10 mM) and NaCl (1.0 M) for 24 h. Although we attempted to soak crystals in buffer solution containing RB more than 10 mM, the crystals cracked during the incubation. Prior to data collection, crystals were continuously immersed in the precipitant solutions containing 10% and 25% (w/w) glycerol and subsequently frozen in liquid nitrogen. X-ray diffraction datasets for the sample were collected at 100 K at beamline BL26B1 at

SPring-8 using X-ray wavelengths of 1.0 Å. The data were processed with the program *HKL2000* [43].

The crystal parameters and the data collection statistics are summarized in Table S2. The structures were solved by molecular replacement with *MOLREP* [44] by using tetragonal and orthorhombic HEWL structures (PDB ID: 193L and 1BGI, respectively). Refinement of the protein structures was performed using *REFMAC5* [45] in the *CCP4* suite [46]. Rebuilding was performed using *COOT* [47] based on sigma-weighted ($2F_o - F_c$) and ($F_o - F_c$) electron density maps. After rigid-body refinement, the coordination structures of Pt binding sites were determined using anomalous Fourier difference maps and geometric parameters. Water molecules were positioned to fit residual ($F_o - F_c$) density peaks with a lower cut-off of 3σ . The models were subjected to quality analysis during the various refinement stages with omit maps and *RAMPAGE* [48]. The refinement statistics are summarized in Table S1. Atomic coordinates for RB·CL-T-HEWL composite are deposited in the Protein Data Bank under accession number: 5YKY.

2.3. Preparation of the composite of CL-X-HEWL ($X = T, O$) with RB (RB·CL-X-HEWL)

HEWL crystals in the tetragonal form were obtained using a batch crystallization method followed by a cross-linking treatment according to the previously reported procedure (CL-T-HEWL) [41]. CL-T-HEWL (100 mg) was soaked in HEPES buffer (10 mM, pH 7.8, 50 mL) containing rose bengal (RB, 1.0 mM) at room temperature for 12 h to immobilize RB in CL-T-HEWL. The composite of CL-T-HEWL with RB was obtained as dark red crystals (RB·CL-T-HEWL). The same procedure was applied to obtain the composite of cross-linked orthorhombic HEWL crystals with RB (RB·CL-O-HEWL) [42].

2.4. Photocatalytic hydrogen evolution by the composite of RB·CL-T-HEWL with Pt nanoparticles via photodeposition method

A phosphate buffer suspension (50 mM, pH 8.0, 2.0 mL) containing RB·CL-T-HEWL (2.5 mg), β -dihyronicotinamide adenine dinucleotide (NADH, 10 mM) and $\text{H}_2\text{PtCl}_6 \cdot 6\text{H}_2\text{O}$ (0.097 mM) was flushed with N_2 gas for 40 min. The suspension was then photoirradiated for a certain time using a white LED light (130 mW, 33.2 mW cm^{-2} , Rely On Ltd., Japan) positioned perpendicular to a cuvette at a distance of 3.0 cm. While photoirradiation, RB·CL-T-HEWL gradually turned to black to form Pt nanoparticles (PtNPs) inside RB·CL-T-HEWL (PtNP@RB·CL-T-HEWL). Gas in a headspace was analyzed by Shimadzu GC-14B gas chromatography (detector: TCD, column temperature: $80 \text{ }^\circ\text{C}$, column: active carbon with the particle size 60–80 mesh, carrier gas: N_2) to determine the amount of evolved H_2 . After the H_2 evolution reaction, buffer solution containing NADH (200 mM, 100 μL) was added to the cuvette for the next run to evaluate recycling performance of PtNP@RB·CL-T-HEWL.

2.5. Preparation of the composite of CL-T-HEWL with RB and Pt nanoparticles by a chemical reduction method

RB·CL-T-HEWL (50 mg) was soaked in an acetate buffer solution (100 mM, pH 8.0, 5.0 mL) containing $\text{H}_2\text{PtCl}_6 \cdot 6\text{H}_2\text{O}$ (50 mM) for 30 min, and then washed with water to remove unbound Pt complexes. CL-T-HEWL containing Pt complexes were soaked in an acetate buffer (100 mM, pH 4.8, 3.0 mL) containing NaBH_4 (250 mM) for 15 min, then washed with water, and dried *in vacuo* (PtNP@RB·CL-T-HEWL(CR)).

3. Results and discussion

3.1. Preparation and catalysis of composite photocatalysts

RB was immobilized in CL-T-HEWL (RB·CL-T-HEWL) by the equilibrium adsorption [49]. The diffuse reflectance UV–vis absorption spectrum of RB·CL-T-HEWL evidenced absorption of RB (Fig. S1). The amount of adsorbed RB per unit cell of CL-T-HEWL was determined to be 0.64 based on the absorbance change at 548 nm of the RB mother liquor induced by the adding CL-T-HEWL (Fig. S2). This amount is comparable to that determined by X-ray fluorescence (XRF) spectroscopy based on the molar ratio of iodine in RB and sulfur in CL-T-HEWL (0.45, Table S1).

RB·CL-T-HEWL was soaked in a buffer solution containing H_2PtCl_6 (0.097 mM) as the precursor for Pt nanoparticles. The crystal structure of soaked RB·CL-T-HEWL was determined at 1.88 Å resolution (PDB ID: 5YKY; Fig. 1a, 1b and Table S2). Two positions for immobilizing the Pt complexes were determined by the anomalous difference maps. A Pt complex at a Pt_a position is bound to N^ϵ of His15 with the $\text{Pt}_a\text{--N}^\epsilon$ length of 2.5 Å, as reported for HEWL crystals with other Pt-halide complexes (Fig. 1c) [50, 51]. Five Cl^- ions bound to the Pt_a ion fulfill the octahedral coordination. The other Pt complex in the form of $[\text{Pt}_b\text{Cl}_6]^{2-}$ was immobilized in the anion exchange site by the positively charged side chain of Lys1 by electrostatic interaction where the distance between Pt_b and N^α of Lys1 was 3.4 Å (Fig. 1d) [50]. Although unambiguous positions of RB were hardly determined, negatively charged RB is expected to be immobilized with positively charged side chains of Arg5, Arg125 and Arg21 from an adjacent HEWL monomer in the solvent channel (Fig. 1b and S3). The distance between a binding site for a Pt complex and the positively charged areas is enough close for electron transfer (*ca.* 2 nm). Pt complexes were converted into PtNPs by the reduction with photogenerated $\text{RB}^{\cdot-}$ followed by the photocatalytic

reaction, because metal nanoparticles are grown in the region where pre-organized metal complexes are bound as reported previously [14].

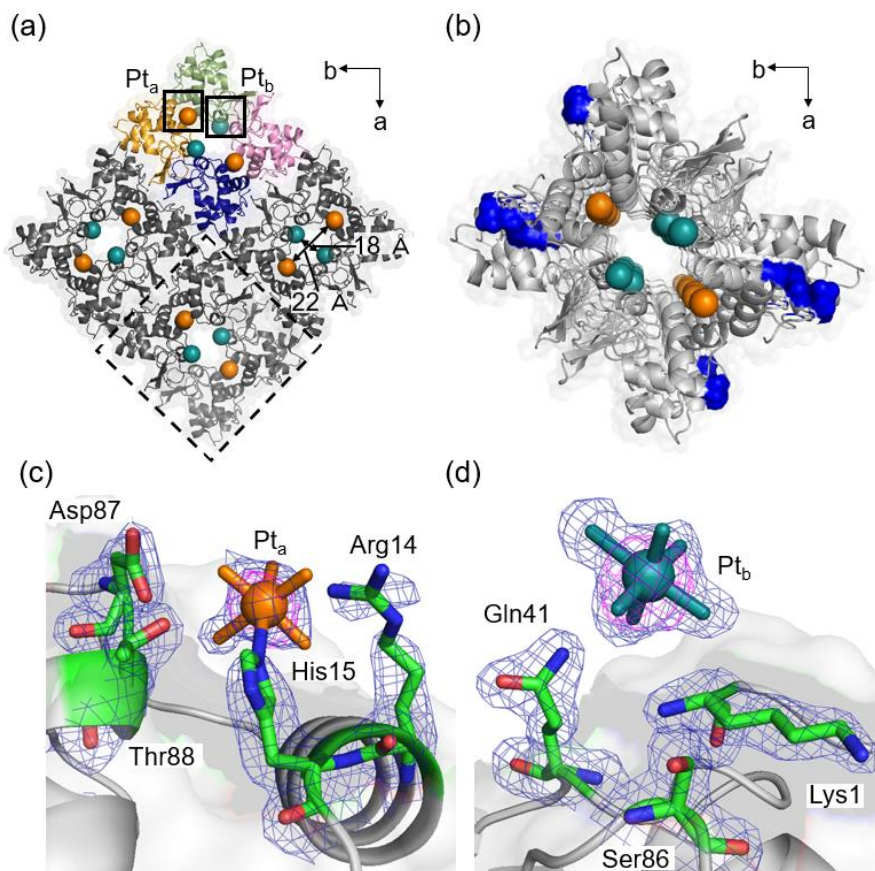


Fig. 1. (a) The overall structure of RB·CL-T-HEWL immobilizing pre-organized Pt-chloride complexes. The Pt atoms are represented by spheres. (b) Magnifications of a pore indicated by a dashed square in (a). Blue residues indicate positively charged area on the surface of solvent channels. (c, d) Magnifications of two Pt binding sites indicated by the black squares in (a). Anomalous difference Fourier maps at 3.0σ indicate the positions of individual Pt atoms (magenta and turquoise). Selected $2F_o - F_c$ electron-density maps at 1.0σ are the deep blue mesh. Crystallographic images were produced by Pymol [52].

Photocatalytic H_2 evolution was performed by the photoirradiation of a phosphate buffer dispersion (pH 8.0) containing RB·CL-T-HEWL (2.5 mg), H_2PtCl_6 (0.097 mM) and NADH (10 mM) as an electron donor as shown in Fig. 2 (solid line). PtNP@RB·CL-T-HEWL exhibited the activity for photocatalytic H_2 evolution with the initial H_2 -evolution rate of $0.73 \mu\text{mol h}^{-1}$ and the

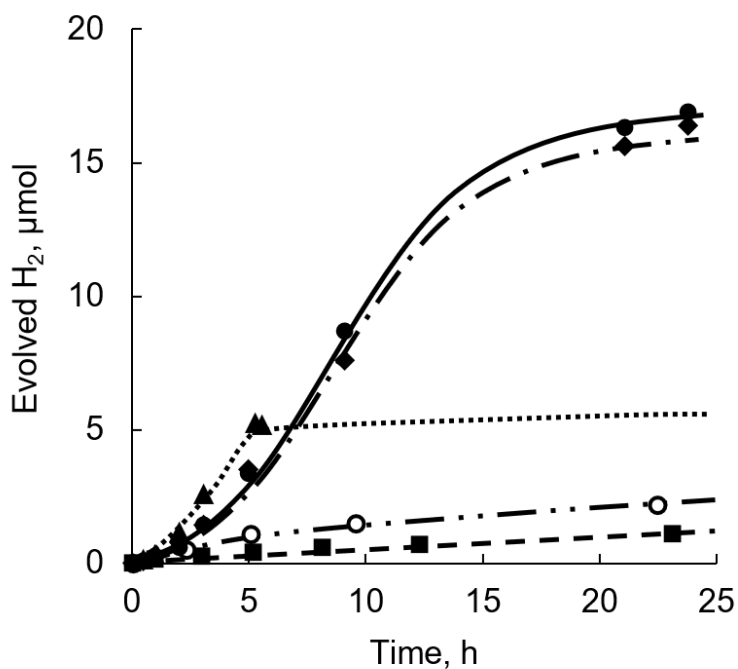


Fig. 2. Time courses of H₂ evolution under photoirradiation ($\lambda > 340$ nm) of a deaerated dispersion of a phosphate buffer (50 mM, pH 8.0, 2.0 mL) containing PtNP@RB·CL-T-HEWL (●, solid line), PtNP@RB·CL-O-HEWL (◆, dashed dotted line), PtNP@RB·CL-T-HEWL(CR) (■, dashed line), homogeneous solution of H₂PtCl₆ with RB (▲, dotted line) and the mixture of homogeneous solution of H₂PtCl₆ with RB and CL-HEWL (○, dashed double-dotted line), in the presence of NADH (10 mM).

H₂ yield of 17 μmol for 24 h, suggesting that a photocatalytic system was successfully constructed in CL-T-HEWL. The H₂ yield based on the used amount of NADH reached 85%, where the turnover frequency (TOF) of evolved H₂ per RB was 13 h⁻¹. The H₂ yield was higher than that (76%) reported for a photocatalytic H₂-evolution system using an assembly of monodispersed silica-alumina nanoparticles as a porous support together with 2-phenyl-4-(1-naphthyl)quinolinium cation, PtNPs and NADH as a photosensitizer, an H₂-evolution catalyst and an electron donor, respectively [36]. Moreover, the TOF was increased from those of catalytic

systems constructed using a spherical mesoporous silica-alumina (0.9 h^{-1}) or a porous assembly of monodispersed silica-alumina nanoparticles (4.6 h^{-1}) as supports [36–38].

A homogeneous reaction system using a buffer solution containing PtNPs, RB and NADH without RB·CL-T-HEWL, which is similar to the method described in the previous report, resulted in lower H_2 yield ($5.3 \text{ } \mu\text{mol}$, 27%, for 6 h) than that with PtNP@RB·CL-T-HEWL although the initial H_2 -evolution rate is relatively high ($0.97 \text{ } \mu\text{mol h}^{-1}$, Fig. 2, dashed line) [53]. The mixture of PtNPs, RB, NADH and CL-HEWL in a buffer solution also showed the low activity for H_2 evolution ($2.2 \text{ } \mu\text{mol}$, 11%, for 23 h, Fig. 2, the dashed double-dotted line) compared with PtNP@RB·CL-T-HEWL. These results suggest that the immobilization of each component in a solvent channel of HEWL crystals enhanced the stability of the catalytic system. Slightly slower H_2 evolution with CL-T-HEWL can be explained by slow diffusion of NADH in the solvent channels. The amounts of RB and Pt per HEWL monomer for PtNP@RB·CL-T-HEWL determined by XRF were 0.035 and 0.036, respectively (Table S1). Formation and growth of PtNPs during the reaction were confirmed by a transmission electron microscope (TEM). PtNPs in the size of *ca.* 1 nm (Fig. S5a) on RB·CL-T-HEWL were observed soon after the induction period (~ 5 h). Growing PtNPs with the size of *ca.* 2 nm were observed at the prolonged reaction time of 24 h (Fig. S5b). Formed PtNPs were well-dispersed throughout the crystals of RB·CL-T-HEWL and smaller than the size of solvent channels of CL-T-HEWL.

RB·CL-O-HEWL with PtNPs (PtNP@RB·CL-O-HEWL) also showed the activity for photocatalytic H_2 evolution (Fig. 2, dotted dashed line). Single-crystal X-ray structure analysis of RB·CL-O-HEWL indicated that electron density originated from a Pt ion was found at the Pt_a position while this density was too low to fit the models of a Pt-chloride complex (Fig. S4 and Table S2). The positively charged area around Arg5 and Arg125 from an adjacent HEWL

monomer is located with separation distances of *ca.* 2 nm from Pt_a (Fig. S4). Thus, the catalytic activity of PtNP@RB·CL-O-HEWL is comparable to that of PtNP@RB·CL-T-HEWL.

CL-T-HEWL with PtNPs was also prepared by chemical reduction (CR) of Pt complexes by NaBH₄ according to a previously reported procedure (PtNP@RB·CL-T-HEWL(CR)) [23]. The amount of Pt per HEWL monomer determined by XRF is about 5 times higher than that of PtNP@RB·CL-T-HEWL (0.16, Table S1). However, only negligible amount of H₂ (< 1 μmol in 12 h) was evolved with PtNP@RB·CL-T-HEWL(CR) (Fig. 2, dashed line). The TEM image in Fig. S5c showed that the particle size of PtNPs in PtNP@RB·CL-T-HEWL(CR) was as small as *ca.* 3 nm. As previously reported, PtNPs generated by CR originated from pre-organized Pt ions in ~10 neighboring HEWL monomers [23]. Thus, PtNPs were immobilized irrespective of the positions of adsorbed RB. On the other hand, PtNPs generated by the reduction with photogenerated RB^{•-} were immobilized in immediate proximity to adsorbed RB, where efficient electron transfer from photogenerated RB to PtNPs is possible.

3.2. Recycling tests

The H₂ evolution lasted for more than 64 h by the successive addition of an aqueous NADH solution to the reaction solution containing PtNP@RB·CL-T-HEWL after H₂ evolution ceased (Fig. 3). The yield at each run together with the initial H₂-evolution rate was 17 μmol (85%, 0.73 μmol h⁻¹), 11 μmol (55%, 2.3 μmol h⁻¹) and 7 μmol (35%, 0.94 μmol h⁻¹), indicating that the total turnover number exceeded 5.9×10². The induction period observed at the first run resulted from the PtNP formation. The RB/HEWL ratio of PtNP@RB·CL-T-HEWL determined by X-ray fluorescence spectroscopy (Table S1) decreased to 0.020 after the third run from that after the first run (0.035) due to partial leaching of RB. On the other hand, Pt/HEWL ratio after the third

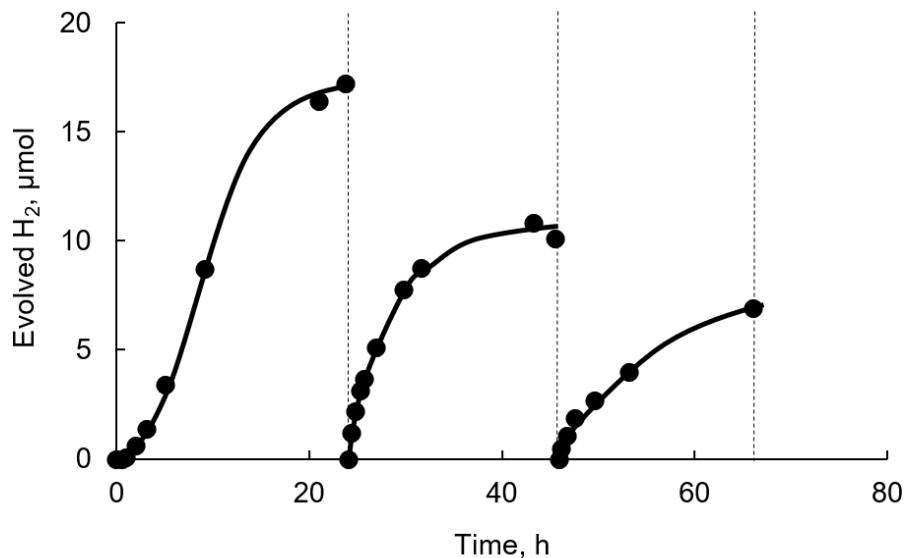


Fig. 3. Time courses of the photocatalytic H₂ evolution under photoirradiation of a deaerated phosphate buffer (50mM, pH 8.0, 2.0 mL) containing H₂PtCl₆ (0.097 mM), RB·CL-T-HEWL (2.5 mg, [RB] = 0.40 mM) and NADH (10 mM) in three repetitive experiments.

run (0.046) increased by 29% based on that after the first run (0.036). Thus, the spontaneous growth of PtNPs which prohibits the diffusion of NADH in the solvent channels and partial leaching of RB are reasons for deceleration of H₂ evolution in repetitive runs.

3.3. Precursor effect for composite photocatalysts

The concentration effects of the Pt precursor in reaction solutions were examined ranging from 0.019 to 0.72 mM (Fig. 4). When the concentration of H₂PtCl₆ was increased to 0.72 mM from 0.097 mM, the initial reaction rate (0.71 μmol h⁻¹, Fig. 4, dashed line) is comparable to that of the original concentration although the amount of evolved H₂ was decreased to almost half (9.0 μmol, 45 %, for 24 h) compared with that of the original concentration. This result suggests that

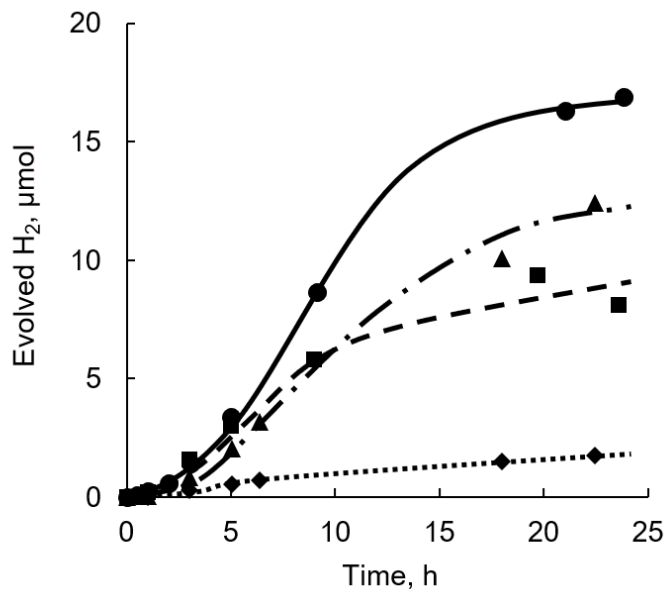


Fig. 4. Time courses of H₂ evolution under photoirradiation (white light) of a phosphate buffer (50 mM, pH 8.0, 2.0 mL) containing NADH (10 mM), RB·CL-T-HEWL (2.5 mg) and various H₂PtCl₆. Concentrations of H₂PtCl₆: 0.019 mM (◆, dotted line), 0.048 mM (▲, dashed dotted line), 0.097 mM (●, solid line) and 0.72 mM (■, dashed line).

the overloaded H₂PtCl₆ formed larger PtNPs, which eventually prohibited the diffusion of NADH in the solvent channels. On the other hand, when the concentrations of H₂PtCl₆ were lower than 0.097 mM, the reaction rate and the amount of evolved H₂ were increased as H₂PtCl₆ increases from (Fig. 4), where no larger PtNPs formed.

Various precursor complexes for PtNPs were employed to scrutinize their effect on the catalysis. Composite catalysts prepared with K₂PtCl₆ and K₂PtCl₄ exhibited high catalytic activity, because PtCl₆²⁻ and PtCl₄²⁻ can bind to His15 and/or to the anion exchange site as reported previously (Fig. 5, solid and dotted lines) [23,54]. A composite catalyst prepared with cationic Pt(NH₃)₄Cl₂ exhibited no activity for photocatalytic H₂ evolution (< 1 μmol in 5 h, Fig. 5, dashed line), resulting from no immobilization of Pt ions in CL-T-HEWL as evidenced by XRF measurements. Strong basicity of the NH₃ ligands and electrostatic repulsion between an

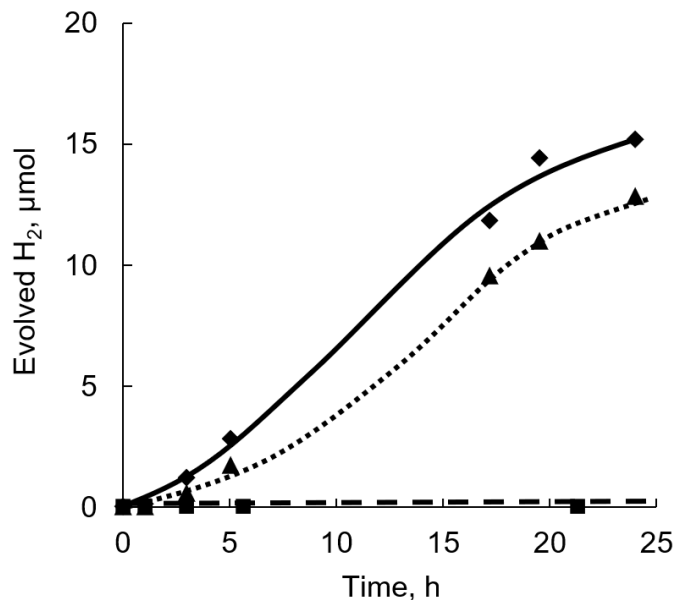


Fig. 5. Time courses of the photocatalytic H₂ evolution under photoirradiation (white light) of a deaerated phosphate buffer (50 mM, pH 8.0, 2.0 mL) containing NADH (10 mM), RB·CL-T-HEWL (2.5 mg) and various PtNP precursors (0.097 mM). A precursor of PtNPs was chosen from K₂PtCl₆, (◆, solid line), K₂PtCl₄, (▲, dotted line), and Pt(NH₃)₄Cl₂ (■, dashed line).

anion exchange site and [Pt(NH₃)₄]²⁺ disturbed [Pt(NH₃)₄]²⁺ from immobilization with neither histidine nor an anion exchange site.

4. Conclusion

We have successfully constructed a heterogeneous catalyst for photocatalytic H₂ evolution using cross-linked HEWL crystals as a porous support. Rose bengal was immobilized in the solvent channels of HEWL crystals and H₂-evolution catalysts, PtNPs, were obtained from pre-organized Pt complexes in immediate proximity to rose bengal by the *in-situ* reduction of Pt complexes. The photosensitizer and the catalytic PtNPs were accumulated in the immediate proximity by utilizing coordination bonds and electrostatic interaction, resulting in efficient H₂ evolution. More sophisticated catalytic systems can be provided by employing cross-linked porous

protein crystals as supports where each component can be precisely arranged by molecular interactions.

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Appendix A. Supplementary content

Supplementary contents associated with this article can be found, in the online version, at <http://dx.doi.org/10.1016/j.cattod. XXXXX>.

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